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U.S. DEPARTMENT OF COMMERCE / National Bureau of Standards

Physical Properties Data Compilations Relevant to Energy Storage.

I. Molten Salts: Eutectic Data

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National Bureau of Standards
MAY 24 1978

Physical Properties Data Compilations Relevant to Energy Storage.

I. Molten Salts: Eutectic Data

George J. Janz, Carolyn B. Allen,
Joseph R. Downey, Jr., and R. P. T. Tomkins

Molten Salts Data Center,
Cogswell Laboratory,
Rensselaer Polytechnic Institute,
Troy, New York 12181

U.S. DEPARTMENT OF COMMERCE, Juanita M. Kreps, Secretary

Dr. Sidney Harman, Under Secretary
Jordan J. Baruch, Assistant Secretary for Science and Technology

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Foreword

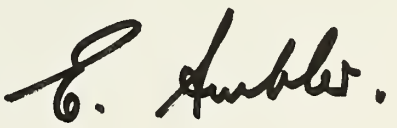
The National Standard Reference Data System provides access to the quantitative data of physical science, critically evaluated and compiled for convenience and readily accessible through a variety of distribution channels. The System was established in 1963 by action of the President's Office of Science and Technology and the Federal Council for Science and Technology, and responsibility to administer it was assigned to the National Bureau of Standards.

NSRDS receives advice and planning assistance from a Review Committee of the National Research Council of the National Academy of Sciences-National Academy of Engineering. A number of Advisory Panels, each concerned with a single technical area, meet regularly to examine major portions of the program, assign relative priorities, and identify specific key problems in need of further attention. For selected specific topics, the Advisory Panels sponsor subpanels which make detailed studies of users' needs, the present state of knowledge, and existing data resources as a basis for recommending one or more data compilation activities. This assembly of advisory services contributes greatly to the guidance of NSRDS activities.

The System now includes a complex of data centers and other activities in academic institutions and other laboratories. Components of the NSRDS produce compilations of critically evaluated data, reviews of the state of quantitative knowledge in specialized areas, and computations of useful functions derived from standard reference data. The centers and projects also establish criteria for evaluation and compilation of data and recommend improvements in experimental techniques. They are normally associated with research in the relevant field.

The technical scope of NSRDS is indicated by the categories of projects active or being planned: nuclear properties, atomic and molecular properties, solid state properties, thermodynamic and transport properties, chemical kinetics, and colloid and surface properties.

Reliable data on the properties of matter and materials are a major foundation of scientific and technical progress. Such important activities as basic scientific research, industrial quality control, development of new materials for building and other technologies, measuring and correcting environmental pollution depend on quality reference data. In NSRDS, the Bureau's responsibility to support American science, industry, and commerce is vitally fulfilled.



Ernest Ambler, *Director*

Preface

This series of publications is aimed at providing physical properties data on materials used in energy storage systems. It was inspired by a requirement in the Department of Energy's Division of Energy Storage Systems for materials property data needed by its contractors in the timely development of energy storage devices. As prime contractor for this program, the Lawrence Livermore Laboratory (LLL) has requested the Office of Standard Reference Data (OSRD) to manage the task of gathering the data, using its established network of data centers and other identified sources of expertise. The OSRD monitors the progress of work, reviews the results, and conveys the numerical data to LLL where the data are converted for entry into an automated data storage and retrieval system. Every effort is made to supply data which have been critically examined in light of the latest knowledge concerning theory and experiment. However it must be recognized that in a rapidly moving technology some of the data will be superseded rather quickly as new materials and techniques are introduced. Thus access to the data via computer terminal as well as publication in this series should help provide the practitioner with timely and useful data which he requires to solve his problems in energy storage.

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Physical Properties Data Compilations Relevant to Energy Storage

I. Molten Salts: Eutectic Data

George J. Janz, Carolyn B. Allen, Joseph R. Downey, Jr., and R. P. T. Tomkins

Molten Salts Data Center, Cogswell Laboratory, Rensselaer Polytechnic Institute, Troy, New York 12181

The present compilation provides an authoritative compendium of melting points, and compositions of molten salt eutectic mixtures. Data for mixtures melting in the range −138°C to 2800°C are reported. Value judgments have not been attempted. Titles of the articles in the literature citations and a system index are included for approximately 6000 eutectic entries.

Key words: Data compilation; eutectic: eutectic composition; eutectic data; eutectic temperature; inorganic compound; melting point; molten salt; phase diagrams.

Introduction

An analysis of energy-related research and development programs shows the emergence of thermal energy storage, advanced batteries, and coal gasification as well-defined areas in the past decade. The potential of inorganic compounds and their mixtures in these technologies is receiving consideration from various practical viewpoints, for example, the relatively large latent enthalpies that accompany the process of melting have directed consideration of such materials in the design of a series of "second-generation" thermal energy storage subsystems of considerably greater capacity than the systems utilizing the storage capacity of fluids (such as water) and rocks. As molten salts, the features of large liquid state ranges, low vapor pressures, high electrical conductivities, and good viscosities are important considerations relative to diverse applications, such as heat transfer fluids, solvents, reaction media, and as molten electrolytes.

While the accumulated store of information on molten salts is considerable, and while several authoritative molten salts treatises have appeared in the past two decades [1–7] [1] a basic difficulty to the potential utilization of such materials is the diversity of the information sources.

The present communication reports the results of a compilation of data for eutectic systems (melting points; compositions), undertaken to provide an authoritative reference for such systems. A partial list of some of the more useful compilations of phase diagrams would include International Critical Tables [8], Landolt-Bornstein [9] Clark [10], Robertson [11], Thoma [12], Voskresenskaya [13], Sinistri et al. [14], Shaffer [15], Franzozini et al. [16], Torapov et al. [17], and Levin et al. [18]. Such sources, as well as the primary resarch literature, were used to develop the summary reported here.

In the compilation of eutectic data, a listing by increasing melting temperatures has been adopted. The temperature range covered is from −138°C to 2800°C. The composition data are reported as mole percent. Value judgments have not been attempted, and where two or more sets of results were reported, these are listed as individual entries. While it was the intent to limit this compilation to inorganics, some organic compounds appeared particularly relevant and these have been included. To enable the location of the data entries by the materials (i.e., salts), both a systems index and a compound index were developed and are included. In the citation of the primary literature, titles have been generally included as part of the bibliography as a further aid on the matter of data origin.

It should be noted that this compilation is an effort to provide an authoritative and comprehensive collection of eutectic data to June 1976. Considerations of safety and hazards, and corrosion and containment are important factors, but fall outside the scope of the present communication. An assessment of these factors for inorganic compounds and their mixtures as molten salts has been completed and has been published elsewhere [19].

[1] Figures in brackets indicate literature references for the introduction.

Acknowledgments

We found the significant contribution by Paul V. Clark (Albuquerque, N.M.) a particularly helpful staging area; without this, our task would have been much more difficult.

The completion of this work owes much to the participation of Mary Leu Waldron, Elizabeth Schoonmaker, Mark J. Chudzicki, Ralph G. Kornbrekke, Timothy P. Walker, Keith M. Smith, and Katherine Petrarca at the Molten Salts Data Center—Rensselaer Polytechnic Institute. We acknowledge this with pleasure.

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Eutectic Data

Table 1 reports data on inorganic eutectics current to June 1976. Also included are systems which form a series of solid solutions with minimum melting point; some solid solutions systems with no minimum melting point are included to provide reference to a system which may be under consideration. Some organic systems are also included, notably the formates and acetates, but coverage is not as complete as for the inorganic systems.

The table is arranged in order of increasing melting point; systems which have no eutectic melting point, e.g. solid solutions without minimum, are listed at the end of the table. It should be kept in mind that a specific eutectic for a given system may not be the one with the lowest melting point. In order to locate the minimum melting eutectic for a given system (or to determine whether a given system forms a eutectic) reference must be made to the System Index.

No attempt at value judgments has been made in the course of this compilation. All available data are included even where conflicts are evident. The reader is left to establish his own "best" values for the eutectic information and the systems index should be useful for this task.

The comments column refers to the melting point column in degrees celsius. In addition several abbreviations have been used in the table and are listed herewith.

APP	approximate
NA	not available
MIN	minimum
LT	less than
GT	greater than

TABLE 1. Eutectic data

Locator number	System	Mol %	T, °C	References
1	BF_3–N_2O	76.6	−138.0 ±.25	575
2	BF_3–SO_2	95.2	−128.6 ±.25	575
3	BCl_3–$GeCl_4$	76	−116.0	2732
4	BCl_3–PCl_3	94	−110.0	2179
5	BCl_3–PCl_3	20	−99.0	2179
6	BF_3–SO_2	38.	−97.1 ±.25	575
7	$TiCl_4$–$VOCl_3$	29	−95.0	2008
8	PCl_3–$TeCl_4$	100 APP	−91.0	2821
9	$TiCl_4$–$VOCl_3$	18	−88.0	1240
10	$TiCl_4$–$VOCl_3$	18.6	−88.0	659
11	VCl_4–$VOCl_3$	27.8	−86.5	659
12	NH_3–$NaBH_4$	94.7	−84.0	1287
13	$CO(NH_2)_2$–NH_3	1.8	−81.5	2033
14	$POCl_3$–$VOCl_3$	3	−80.8	1240
15	BBr_3–$SiCl_4$	26.4	−80.0	971
16	NH_3–$NH_4Br(NH_3)_4$	93	−80.0	1005
17	NH_3–NH_4NO_3	96	−79.5	2033
18	KNH_2–NH_3	99 APP	−78.8	961
19	BBr_3–$GeCl_4$	40	−74.0	971
20	$POCl_3$–$SiCl_4$	4 APP	−71.8	1240
21	$SiCl_4$–WCl_6	100 APP	−70.0	2451
22	$SiCl_4$–$TeCl_4$	99 APP	−68.0	3053
23	$FeCl_3$–$GeCl_4$–$TeCl_4$	NA	−59.0 APP	2849
24	BBr_3–$SiBr_4$	76	−57.5	971
25	$AsBr_3$–S_2Br_2	13.5	−56.0	453 1081
26	$AsBr_5$–BBr_3	5.8	−54.1	342
27	$GaCl_3$–$GeCl_4$	7 APP	−54.0	1874
28	BBr_3–$GeBr_4$	80.9	−53.0	971
29	BBr_3–$SnBr_4$	81.7	−52.5	342
30	$GeCl_4$–$SnCl_4$	91 SER SOLID SOL	−51.3	2732
31	$GeCl_4$–$TiCl_4$	96.7 SER SOLID SOL	−50.8	2732
32	Si_2OCl_6–$TiCl_4$	42	−50.0	2008
33	PbR_3–$SnBr_4$	84	−50.0	453 1081
34	CsF–HF	17.2	−49.5	566
35	$SbCl_5$–$TiCl_4$	35	−47.5	1679
36	S_2Br_2–$SnBr_4$	95.5	−47.5	1081
37	$AlBr_3$–BBr_3	5.24	−46.1	342
38	$POCl_3$–$SnCl_4$	5 APP	−33.8	2359 2555
39	$FeCl_3$–$SnCl_4$–$TeCl_4$	NA	−33.0 APP	2849
40	$SnCl_4$–WCl_6	100 APP	−32.0	2451
41	$AsBr_3$–Br_2	34	−31.5	1081
42	$AlCl_3$–$SnCl_4$	.45 LT	−30.5 APP	2359
43	$AsCl_3$–$SbCl_3$	97 APP	−30.0 APP	1279
44	NH_3–$NaBH_4$	69	−28.0	1287
45	NH_3–$NaBH_4$	56.6	−25.1	1287
46	$GaCl_3$–$TiCl_4$	10 APP	−25.0	1874
47	CCl_4–$GaCl_3$	92 APP	−24.0	1874
48	$POCl_3$–$SbCl_5$–$TiCl_4$	NA	−24.0	1679
49	$TiCl_4$–WCl_6	100 APP	−23.0	2451
50	$Ba(BH_4)_2$–N_2H_4	13	−22.0	1253
51	Br_2–N_2O_4	17	−18.0	1058
52	$GeBr_4$–$POCl_3$	36 APP	−16.0	3083
53	$AsCl_3$–$TeCl_4$	100 APP	−16.0	2821
54	Br_2–$SbBr_3$	81.8	−15.5	1081
55	$AlBr_3$–Br_2	21.4	−13.5	1081
56	$POCl_3$–$ReOCl_4$	97	−10.0	2708
57	$GaCl_3$–$SbCl_5$	18	−4.0	2676
58	$GaCl_3$–$POCl_3$	10	−3.8	2202

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References
59	$BeCl_2$–$POCl_3$	2.5	−3.0 APP	1389	
60	$NbCl_5$–$POCl_3$	4 APP	−2.4	1182	
61	$NbCl_5$–$POCl_3$–$TiCl_4$	3.75–92.5–3.75	−2.2	987	
62	$POCl_3$–$TiBr_4$	NA	−2.0	2970	
63	$POCl_3$–$TaCl_5$–$TiCl_4$	93.8–3.1–3.1	−1.7	987	
64	$POCl_3$–$TiCl_4$	3	−0.6	1240	
65	$AlCl_3$–$POCl_3$	3 LT	−0.2	1159	
66	$POCl_3$–$TeCl_4$	99 APP	−0.1	2691	
67	$POCl_3$–$SbCl_5$–$TiCl_4$	94.6–2.68–2.68	0.0	1679	
68	$POCl_3$–WCl_6	100 APP	1.0	2451	
69	$GaCl_3$–$POCl_3$	70	1.5	2202	
70	$SbCl_5$–WCl_6	100 APP	2.0	2451	
71	$POCl_3$–$TeCl_4$	98	3.0	2821	
72	$AsBr_3$–$SnBr_4$	45	3.5	453	1081
73	ICl–$TeCl_4$	80	10.0	2965	
74	ICl–$TaCl_5$	80	10.0	2965	
75	$POCl_3$–$TiBr_4$	NA	10.0	2970	
76	MoF_6–UF_6	78	13.7	2358	
77	SO_3–SeO_3	99.97	15.0	842	
78	$POCl_3$–$ReOCl_4$	20	16.0	2708	
79	ICl–$SeCl_4$	80	16.0	2965	
80	CsF–HF	29.1	16.9	566	
81	H_2O–KF	78	17.0	3222	
82	H_2O–NaCl–Na_2SO_4	89.9–8.9–1.2	18.0	3222	
83	$AlBr_3$–$SnBr_4$	23	20.0	1081	
84	$AlBr_3$–$SnBr_4$	26	20.0	453	
85	ICl–$NbCl_5$	92	21.0	2965	
86	$CaCl_2$–H_2O–$MgCl_2$	8.1–87.6–4.3	22.0	3222	
87	H_2O–K_2CO_3–Na_2CO_3	97.3–1.0–1.7	22.0	3222	
88	$AsBr_3$–PBr_5	82.5	23.5	453	1081
89	$Ca(NO_3)_2 \cdot 4H_2O$–$Zn(NO_3)_2 \cdot 6H_2O$	51	25.0	3221	
90	H_2O–K_2CO_3–Na_2CO_3	86.7–11.8–1.5	25.0	3222	
91	H_2O–$Ni(NO_3)_2$–NH_4NO_3	78.3–13.0–8.7	25.0	3222	
92	$CaCl_2$–H_2O–$MgCl_2$	10.3–86.8–2.9	25.0	3222	
93	$AlBr_3$–$AsBr_3$	26	25.5	453	1081
94	$SbBr_3$–$SnBr_4$	6	27.0	1081	
95	$AlBr_3$–$AsBr_3$	26	28.0	64	
96	$FeCl_3$–$ReOCl_4$	0 APP	29.0	2708	
97	$ReCl_5$–$ReOCl_4$	0 APP	29.0	2708	
98	$ReOCl_4$–$TaCl_5$	100 APP	29.0	2708	
99	$NbCl_5$–$ReOCl_4$	0 APP	29.0	2708	
100	$AlCl_3$–$ReOCl_4$	0 APP	29.0	2708	
101	$MoOCl_4$–$ReOCl_4$	0 APP	29.0	2708	
102	H_2O–$LiNO_3$	73.7	29.0	3222	
103	$GaCl_3$–$SeCl_4$	79 APP	30.0	1856	
104	$GaCl_3$–$ZnCl_2$	68 APP	30.0	1017	
105	$GaCl_3$–PCl_5	70 APP	30.0 APP	2648	
106	XeF_2–XeF_6	.5 APP	30.0 ±10	2952	
107	$Ca(NO_3)_2 \cdot 4H_2O$–$Mg(NO_3)_2 \cdot 6H_2O$	69	30.0	3221	
108	H_2O–K_2CO_3–Na_2CO_3	87.0–10.9–2.1	30.0	3222	
109	$Mg(NO_3)_2 \cdot 6H_2O$–$Zn(NO_3)_2 \cdot 6H_2O$	20	32.0	3221	
110	BiI_3–SiI_4	70	33.8 ±1	2127	
111	$CO(NH_2)_2$–$NaNO_3$–NH_4NO_3	51–8–41	35.0	2038	
112	$Al(NO_3)_3 \cdot 9H_2O$–$Ca(NO_3)_2 \cdot 4H_2O$	20	35.0	3221	

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References				
113	$CsF-HF$	36.1	38.3	566					
114	$SbCl_3-SbI_3$	81.8	41.5	1918					
115	$CaCl_2-H_2O-MgCl_2$	15.8–83.0–1.2	42.0	3222					
116	$CO(NH_2)_2-KNO_3-NH_4NO_3$	48–5–47	43.0	2038					
117	$RbCl-SbCl_3$	15	44.0	1133					
118	$CO(NH_2)_2-NH_4Cl-NH_4NO_3$	52.1–3.8–44	44.0	1746					
119	$Ba(BH_4)_2-N_2H_4$	41.9	44.0	1253					
120	$CO(NH_2)_2-NH_4NO_3$	59.6	44.5	2033					
121	$GaCl_3-TeCl_4$	80 APP	45.0	1856					
122	$GaCl_3-SbCl_3$	73.0	46.4	2621					
123	SbF_5-XeF_2	62.5	50.0	2283					
124	$CO(NH_2)_2-KI-NaI$	79–3–18	50.0	1477					
125	$AlCl_3-GaCl_2-GaCl_3$	12–38–50	50.0	2629					
126	$AlCl_3-BaCl_2-NaCl$	63.5–2.5–34	50.0	3034					
127	$GaCl_3-SbCl_3$	33.2	50.6	2621					
128	$CdCl_2-GaCl_3$	5 APP	52.0	1017					
129	$H_2O-Mg(NO_3)_2-NH_4NO_3$	66.9–10.8–22.3	52.0	3222					
130	$H_2O-Mg(NO_3)_2$	80.1	53.0	3222					
131	$AlBr_3-SbCl_3$	14.5	54.0	1080					
132	$GaCl_3-NaCl$	75	55.0	1231					
133	$CsCl-GaCl_3$	3	55.0	1016					
134	$GaCl_3-HgCl_2$	92	55.0	1017					
135	$SbBr_3-SbCl_3$	70 APP	55.0	1918					
136	$AlCl_3-NaI \cdot AlI_3$	45 APP	55.0	8397					
137	$KCHO_2-KCNS-KNO_3$	48.3–34.9–16.8	55.5	2712					
138	$GaCl_3-LiCl$	82	57.0	1016					
139	$KCl-SbCl_3$	24 APP	57.0 APP	1918					
140	$GaCl_3-NH_4Cl$	85 APP	58.0	1016					
141	$CO(NH_2)_2-NaI$	80.5	58.0	1100	1477				
142	$AlBr_3-SbCl_3$	75.8	59.0	1080					
143	$AlCl_3-HgBr_2$	42.2	59.0	1080					
144	$CO(NH_2)_2-NaI-NaNO_3$	80–18–2	59.0	1100					
145	$GaCl_3-KCl-MgCl_2$	89–3–3	59.0	2613					
146	$CuCl-GaCl_3$	5.2	60.0	905					
147	$AlBr_3-InBr_3$	90	60.0	2888					
148	$GaBr_3-SbBr_3$	46	60.8 METASTABLE	2621					
149	$Al(NO_3)_3 \cdot 9H_2O-Mg(NO_3)_2 \cdot 6H_2O$	38	61.0	3221					
150	$GaCl_3-NaCl$	97	62.0	905					
151	SbF_5-XeF_2	80.5	63.0	2283					
152	$GaCl_3-KCl-MgCl_2$	92–3–5	64.0	2613					
153	$BeCl_2-GaCl_3$	5.7	64.5	1290					
154	$GaCl_3-KCl$	84	65.0	1016					
155	$CO(NH_2)_2-KBr-NaBr$	79–1.5–19.5	65.0	1477					
156	$CO(NH_2)_2-NaBr$	79.5	66.0	1100	1477				
157	$SeCl_4-SbCl_3$	6	66.0	2619					
158	$GaCl_2-GaCl_3$	41	66.0	2629					
159	$GaCl_3-MoCl_5$	93 APP	67.0	2110					
160	$CO(NH_2)_2-Sr(NO_3)_2$	85	67.0	1749					
161	$CsCl-SbCl_3$	7.5	68.0	1133					
162	$CsNO_3-Pb(NO_3)_2$	68	68.0	1082					
163	$Ba(NO_3)_2-CO(NH_2)_2-LiNO_3$	1–83–16	68.0	993					
164	$Ba(NO_3)_2-CO(NH_2)_2-NaNO_3$	4–78.5–17.5	68.0	993					
165	$GaBr_3-SbBr_3$	37.2	68.2	2621					
166	$AlCl_3-GaCl_2-GaCl_3$	14–49–37	69.0	2629					
167	$AlCl_3-KCl-NaCl$	66–14–20	70.0	34	44	79	80	688	
168	$AlCl_3-MoCl_5-NaCl$	63–4–33	70.0 ±3	912					
169	$GaCl_3-KCl$	82.5	70.0 APP	2373					

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References
170	$AlCl_3$–$SbCl_3$	7.5	70.0	1918	
171	$Cd(NO_3)_2$–$NaNO_3$–NH_4NO_3	44.1–.7–55.1	70.0	1031	
172	$AlCl_3$–$CsCl$–$GaCl_3$	NA	70.0	2724	
173	$BiCl_3$–$GaCl_3$	9	70.0	2964	
174	H_2O–LiI	71.2	70.0	3222	
175	$Cd(NO_3)_2$–$NaNO_3$–NH_4NO_3	33.6–2.4–64	71.0	1031	
176	$GaCl_3$–$HgCl_2$	60	72.0	1017	
177	$CO(NH_2)_2$–H_3PO_4	71.5	72.0	2226	
178	H_2O–KCl–$MgCl_2$–$MgSO_4$–$NaCl$	89.9–1.7–7.4–.4–.6	72.0	3222	
179	$AlBr_3$–$SbBr_3$	31	72.2	1918	
180	$GaBr_3$–$SbBr_3$	52.2	72.2	2621	
181	$SbCl_3$–WCl_6	100 APP	73.0	2054	
182	$SbCl_3$–$WOCl_4$	98.8	73.0	2054	
183	$AlBr_3$–$AlCl_3$	65	73.0	64	
184	$CO(NH_2)_2$–KNO_3–$NaNO_3$	70.7–7–22.3	73.0	2389	
185	$CsC_2H_3O_2$–$KC_2H_3O_2$–$LiC_2H_3O_2$	52.5–27.5–20	73.0	1464	
186	SbF_5–XeF_2	57.0	73.5	2283	
187	$GaCl_3$–$MgCl_2$	98.85	73.6	1290	
188	$AlCl_3$–$GaCl_3$	6	74.0	248	
189	$Cd(NO_3)_2$–NH_4NO_3	23.5	74.0	1031	
190	$AlCl_3$–$GaCl_3$	6	74.0	2629	
191	$AlCl_3$–$GaCl_3$	5.7	74.7	1341	
192	$FeCl_3$–$GaCl_3$	0.7	74.7	1296	
193	$AlBr_3$–$SbBr_3$	70.7	74.8	1918	
194	$AlCl_3$–$HfCl_4$–KCl	68.8–3.8–27.4	75.0	1124	
195	$GaCl_3$–$InCl_3$	98.7 APP	75.0 APP	1056	
196	$POCl_3$–$TiBr_4$	NA	75.0	2970	
197	$GaCl_3$–$InCl_3$	99.75	75.9	1341	
198	$AlCl_3$–$HfCl_4$–$NaCl$	60.5–1.6–38	76.0	1124	
199	$AgCl$–$GaCl_3$	6.1	76.0	905	
200	$BaCl_2$–$GaCl_3$	.75	76.0	1290	
201	$CoCl_2$–$GaCl_3$	0.3	76.0	2058	
202	$GaCl_3$–$MnCl_2$	99.7	76.0	2058	
203	AgI–$AgIO_3$–$AgNO_3$	47–18–35	76.0	1094	
204	$CO(NH_2)_2$–KNO_3–$NaNO_3$	NA	76.0	993	
205	$CO(NH_2)_2$–$LiNO_3$	83.5	76.0	993	1009
206	$CaCl_2$–$GaCl_3$	0.7	76.1	1290	
207	H_2O–LiI	37.2	77.0	3222	
208	$GaCl_3$–$NbCl_5$	100 APP	77.9 APP	2110	
209	$AlCl_3$–$NaCl$–$TaCl_5$	63.1–37.2–1.52	78.0	331	
210	$CO(NH_2)_2$–$NaCl$–$NaNO_3$	79–4–17	78.0	1106	
211	$LiNO_3$–NH_4NO_3	25.3 METASTABLE	79.5	3157	
212	$Ba(NO_3)_2$–$CO(NH_2)_2$–KNO_3	5.5–86–8.5	80.0	993	
213	$NaCl{\cdot}AlCl_3$–$NaI{\cdot}AlI_3$	60 APP	80.0	3237	
214	$GaBr_3$–$RbBr$	80	80.0	2809	
215	AgI–$AgNO_3$	55	80.0	3168	
216	$AgNO_3$–HgI_2	81	81.0 APP	1267	
217	S_2Cl_2–$SeCl_2$	100 APP	81.5	2639	
218	$AgNO_3$–$TlNO_3$	48.6	82.8	1943	
219	$AgNO_3$–$TlNO_3$	51.5	82.8	1943	
220	$AlBr_3$–$RbBr$	74	83.0	2470	
221	$CO(NH_2)_2$–$NaNO_3$	77.5	83.0	2389	
222	H_2O–KCl–$MgCl_2$–$MgSO_4$–$NaCl$	89.4–2.2–7.2–.4–.8	83.0	3222	
223	H_2O–SrI_2	85	83.0	3222	
224	S_2Cl_2–$TeCl_4$	100 APP	83.5	2639	
225	XeF_2–XeF_4	35 APP	83.5	2979	
226	BCl_3–$POCl_3$	50	83.8	795	
227	$NbCl_5$–$POCl_3$–$TiCl_4$	21.7–50–28.3	83.8	987	

Table 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
228	$CO(NH_2)_2$–$NaNO_3$	77	84.0	993 1100
229	$AlCl_3$–$GaCl_2$	30	84.0	2629
230	$SbBr_3$–$TeBr_4$	96	84.0	2841
231	$SbBr_3$–SbI_3	85	84.5 ±.5	1918
232	$AlCl_3$–KCl–LiCl	56–7–37	84.5 ±0.5	2975
233	$POCl_3$–$TaCl_5$–$TiCl_4$	50–21–29	86.0	987
234	$AlCl_3$–$SbBr_3$	8.4	86.0	1080
235	$GaBr_3$–$HgBr_2$	67	86.0	2911
236	XeF_2–XeF_4	62 APP	87.0	2979
237	AgI–Ag_2SO_4–Tl_2SO_4	53%SO4,97%AG	87.0	3117
238	$NbCl_5$–$POCl_3$–$TiCl_4$	11.1–70–25.9	87.8	987
239	$AlCl_3$–$FeCl_3$–$MoCl_5$	NA	88.0	896
240	$AlBr_3$–KBr	74	88.0	40 270 688
241	$CO(NH_2)_2$–KI	82.5	88.0	1477
242	$GaCl_3$–KCl–$MgCl_2$	61–37–2	88.0	2613
243	Ga–$GaBr_3$	20	88.0	3018
244	$AlCl_3$–NaCl–$TeCl_4$	60–22–18	88.0	2834
245	$AlCl_3$–KCl–NaCl	63.5–16.5–20	88.5 ±.5	34 44 79 80 688
246	$NbCl_5$–$POCl_3$–$TiCl_4$	13.3–56.6–30.1	89.0	987
247	$AlBr_3$–KBr	75	89.0	2265
248	$Ca(NO_3)_2$-$CO(NH_2)_2$–KNO_3	8.5–88.7–2.8	89.0	2117
249	$AlCl_3$–NaCl–WCl_6	58–40–2	90.0 ±2	1106
250	AlI_3–NH_4I	57 APP	90.0 APP	2284
251	$CO(NH_2)_2$–$H_4P_2O_7$	80	90.0	1799 2226
252	$CsC_2H_3O_2$–$KC_2H_3O_2$–$NaC_2H_3O_2$	50–24–24	90.0	2568
253	$LiNO_3$–NH_4NO_3	26.6	90.4	3157
254	H_2O–$KAl(SO_4)_2$	91.6	91.0	3222
255	$NbCl_5$–$POCl_3$–$TiCl_4$	27.1–57.6–15.3	91.2	987
256	$Ba(NO_3)_2$-$CO(NH_2)_2$–$LiNO_3$	1–55–44	92.0	993
257	$Ca(NO_3)_2$-$CO(NH_2)_2$	9.3	92.0	2109
258	HF–XeF_2	NA	92.0	2789
259	$CO(NH_2)_2$–$Ca(NO_3)_2$	81.9	92.0	3222
260	$Ba(NO_3)_2$-$CO(NH_2)_2$–$LiNO_3$	1.0–55.0–44.0	92.0	3222
261	$POCl_3$–$TaCl_5$–$TiCl_4$	54.5–9.1–36.4	92.4	987
262	$AlCl_3$–KCl–NaCl	60–14–26	93.0	34 44 79 80 688
263	$AlCl_3$–NaCl	66	93.0	27 34 44 45 78 80 84 258 331 451 472 688
264	$AgNO_3$–HgI_2	58	93.0	1267
265	$LiNO_2$–$Sr(NO_2)_2$–$TlNO_2$	24.5–6.4–69.1	93.0	1853
266	$LiNO_2$–$TlNO_2$–$TlNO_3$	45–10–45	93.0	1974
267	$CaCl_2$–H_2O–$MgCl_2$	16.7–80.0–2.3	93.0	3222
268	$AlCl_3$–KCl–NaCl	62.13–12.7–25.17	94.0	34 44 79 80 688
269	$LiNO_2$–$CsNO_2$–$Sr(NO_3)_2$	35.5–62.9–1.5	94.0	1899
270	KNO_3–$LiNO_3$–$TlNO_3$	34–33–33	94.0	144
271	$POCl_3$–$TiCl_4$	59	94.7	2555
272	$AlCl_3$–$BeCl_2$	48	95.0 APP	1060
273	$AlCl_3$–$POCl_3$–$TiCl_4$	8.5–65.3–26.2	95.0	967
274	$AlBr_3$–NaBr	82	95.0	40 688
275	$Cd(NO_3)_2$–$TlNO_3$	27.4	95.0	2181
276	$CO(NH_2)_2$–$LiNO_3$	54.5	95.0	993 1009
277	CsBr–$GaBr_3$	20	95.0	2809
278	GaI_3–SbI_3	45	95.3 METASTABLE	2621
279	$POCl_3$–$TaCl_5$–$TiCl_4$	64.2–7.5–28.3	96.0	987
280	$POCl_3$–$TiCl_4$	58	96.0	1240
281	$Ba(NO_3)_2$-$CO(NH_2)_2$	10.5	96.0	993
282	$GaBr_3$–KBr	71	96.8	3018
283	$AlCl_3$–NaCl–$NbCl_5$	61–37–2	97.0	331
284	AlI_3–KI	69	97.0	1918

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References	
285	$CsNO_3$–KNO_3–$LiNO_3$	24–39–37	97.0	2275		
286	$AlCl_3$–$POCl_3$–$TiCl_4$	7.5–50–42.5	97.8	967		
287	$GaCl_3$–KCl	64	98.0	1016		
288	$POCl_3$–$SbCl_5$–$TiCl_4$	65.3–4.17–30.55	98.0	1679		
289	KNO_2–$LiNO_2$–$Sr(NO_2)_2$	61.2–36.7–2	98.0	2285		
290	AgI–$AgNO_3$–$NaNO_3$	59–40–1 APP	98.0	3115		
291	$AlCl_3$–$POCl_3$–$TiCl_4$	7.5–51.9–40.6	98.2	967		
292	$AlCl_3$–$NbOCl_3$	61.8	100.0	2565		
293	$AlCl_3$–$SeCl_4$	65.2	100.0	50		
294	AlI_3–KI	70 APP	100.0 APP	2284		
295	InI_3–KI	85	100.0	1970		
296	HgI_2–InI_3	75	100.0	1877		
297	$AgNO_3$–AgI	47.9	100.0	228		
298	$LiNO_2$–$TlNO_2$	25	100.0	1148		
299	$CO(NH_2)_2$–NH_4Cl	82.9	101.5	1746		
300	$AgNO_3$–NH_4NO_3	30 APP	101.5	3122		
301	$AlBr_3$–$RbBr$	78.5	102.0	2265		
302	$CsNO_2$–$LiNO_2$	60	102.0	1192		
303	KNO_3–$LiNO_2$	45	102.0	917		
304	$CO(NH_2)_2$–K_2CO_3	93	102.0	2149		
305	$CsNO_2$–$KC_2H_3O_2$–KNO_2	40.5–39.5–20	102.0	2908		
306	AgI–$AgNO_3$–$NaNO_3$	43.5–55–1.5	102.0	3115		
307	$AgNO_3$–NH_4NO_3	30.9	102.4	3123		
308	$POCl_3$–$TaCl_5$–$TiCl_4$	56.3–31–12.7	102.6	987		
309	$POCl_3$–$SbCl_5$–$TiCl_4$	51.2–2.33–46.51	102.7	1679		
310	$Cd(NO_3)_2$–KNO_3–$NaNO_3$	46–14.6–39.4	103.0	2152		
311	$AlCl_3$–$NaNbOCl_4$	78.6	104.0	2086		
312	$AlCl_3$–KBr	65.5	104.0	1080		
313	KNO_2–$LiNO_2$	59.3	104.0	1201		
314	$CsC_2H_3O_2$–$Zn(C_2H_3O_2)_2$	55	104.0	2957		
315	AlI_3–KI	67.5	105.0	2523		
316	$CsNO_2$–$LiNO_2$	39	105.0	1201		
317	$Ba(NO_2)_2$–KNO_2–$LiNO_2$	0.25–59.15–40.6	106.0	2116		
318	$KC_2H_3O_2$–$KCNS$–KNO_3	19.5–54.5–26.0	106.0	3236		
319	$AlCl_3$–$NaCl$–$WOCl_4$	60–40–0 APP	107.0	2467		
320	KNO_2–$LiNO_2$	60	107.0	2285		
321	$Ca(NO_3)_2$–$NaNO_3$–NH_4NO_3	9.8–13.6–76.6	107.5	684		
322	$AlCl_3$–$BiCl_3$–$NaCl$	58–12–30	108.0	2525		
323	$AlCl_3$–$NaCl$	61	108.0	34	45	688
324	$AlCl_3$–$TeCl_4$	62.5	108.0	50		
325	$AgNO_3$–HgI_2	43	108.0	1267		
326	$LiNO_2$–$RbNO_2$–$Sr(NO_2)_2$	47.4–47.4–5.3	108.0	1899		
327	KNO_2–$LiNO_3$	60	108.0	917		
328	$FeCl_3$–$NaCl$–$TeCl_4$	39.5–14.5–46.0	108.0	3021		
329	$CO(NH_2)_2$–KBr	88.5	109.0	1477		
330	$CsNO_2$–$LiNO_2$	59.4	109.0	1201		
331	$CO(NH_2)_2$–KNO_3	85	109.0	2389		
332	$FeCl_3$–KCl–$LiCl$	52–16–32	109.0 APP	2966		
333	$CsC_2H_3O_2$–$CsNO_2$–$KC_2H_3O_2$	29.5–35.5–35	109.0	2908		
334	$FeCl_3$–KCl–$LiCl$	51–17–32	109.5 ±0.5	2975		
335	$AgNO_3$–NH_4NO_3	50 APP	109.6	3122		
336	$AlCl_3$–$NbCl_5$–$TaCl_5$	69.8–17.2–13	110.0	658		
337	AlI_3–NH_4I	67 APP	110.0 APP	2284		
338	GaI_3–SiI_4	75	110.0	2289		
339	InI_3–SnI_2	80	110.0	1345		
340	$Ca(NO_3)_2$–NH_4NO_3	16.9	110.0	1952		
341	$CO(NH_2)_2$–KNO_3	84	110.0	993		
342	$AlCl_3$–AlI_3	40	110.0	2636		

Table 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
343	$AlCl_3$–NaCl–$TeCl_4$	43–16.5–40.5	110.0	2834
344	AgI–$AgNO_3$–$NaNO_3$	26–73–1	110.0	3115
345	$POCl_3$–$SbCl_5$–$TiCl_4$	52.8–41.51–5.66	110.5	1679
346	$Ca(NO_3)_2$–NH_4NO_3	16.6	111.0	684
347	$Cd(NO_3)_2$–KNO_3–$LiNO_3$	11.3–61.2–27.5	111.0	512
348	$FeCl_3$–KCl–LiCl	50–17–33	111.0 APP	2966
349	$FeCl_3$–KCl–LiCl	50–17–33	111.5 ±0.5	2975
350	$AlCl_3$–NaCl	62	112.0	27 34 44 45 78 80 84 258 331 451 472 688
351	$CO(NH_2)_2$–NaCl	90	112.0	1100 1477
352	$CsNO_3$–CsOH	32.5	112.0	1467
353	AgCl–$AgNO_3$–KNO_3	2.8–59.5–37.7	113.0	376
354	$LiNO_2$–$RbNO_2$–$Sr(NO_2)_2$	28.7–68.7–2.6	113.0	1899
355	$CsNO_3$–$LiNO_2$	35	113.0	1192
356	$LiNO_3$–$RbNO_2$	30	113.0	1012
357	$AlCl_3$–$GaAlCl_4$	15	113.0	2629
358	CuCl–$FeCl_3$–$TeCl_4$	5–62–33	113.0	2918
359	AlI_3–SnI_4	40	113.5	1918
360	$AlCl_3$–LiCl	60	114.0	84 688
361	$AlCl_3$–$TeCl_4$	39.5	114.0	50
362	$KCHO_2$–KNO_3	62.1	114.0	2712
363	$NaAlCl_4$–$NbOCl_3$–$TaCl_5$	75–5–20	114.0	3107
364	$GaBr_3$–TlBr	70	114.0	2809
365	$AlCl_3$–$POCl_3$	63.5	114.3	1182 2359
366	$AlCl_3$–$POCl_3$	62–63	114.4	1159
367	$AlCl_3$–NaCl	60	115.0	27 34 44 45 78 80 84 258 331 451 472 688
368	$AlCl_3$–KCl–$NbCl_5$	50–45–5	115.0	332
369	$AlCl_3$–$POCl_3$	65	115.0	540
370	$AlCl_3$–CsCl–$TaCl_5$	69.7–18.7–11.6	115.0	240
371	$CO(NH_2)_2$–KCl	91	115.0	1477
372	$CsNO_2$–$LiNO_2$	40	115.0	1192
373	$CsNO_3$–KNO_3–$LiNO_3$	33–36–31	115.0	2275
374	KNO_3–$NaNO_3$–NH_4NO_3	7.5–18–74.5	115.0	1009
375	$CsC_2H_3O_2$–$NaC_2H_3O_2$	68	115.0	1013
376	AlI_3–CsI–NaI	71.0–5.5–23.5	115.0	2715
377	CuCl–$FeCl_3$–$TeCl_4$	10–48–42	115.0	2918
378	$AgNO_3$–$Cd(NO_3)_2$–KNO_3	26.8–50.2–23	115.0	3119
379	$AlCl_3$–KCl–$ZrCl_4$	65–33–2	116.0	794
380	$AlCl_3$–$KNbOCl_4$	78.4	116.0	2086
381	$AlCl_3$–$ZrCl_4$	83	116.0	794
382	AlI_3–AsI_3	41	116.0	1918
383	InI_3–SiI_4	15	116.0	2682
384	$CsC_2H_3O_2$–$CsNO_2$–$RbC_2H_3O_2$	32.5–35–32.5	116.0	2912
385	$NbCl_5$–$POCl_3$	52.5	116.8	1182
386	$LiNO_2$–$RbNO_2$	35.2	117.0	1201
387	$AlCl_3$–KCl–LiCl	50–13–37	117.0 APP	2975
388	$Ca(NO_3)_2$–KNO_3–$LiNO_3$	8.8–59.1–32.1	117.4	557
389	$AlCl_3$–$TaCl_5$	61.2	118.0	241 331 658
390	BiOI–InI_3	20	118.0	1878
391	GaI_3–GeI_4	74	118.0	2289
392	InI_3–TlI	85	118.0	1970
393	KNO_3–$NaNO_3$–NH_4NO_3	10.3–20.6–69.1	118.5	684
394	KNO_3–$NaNO_3$–NH_4NO_3	9.4–20.0–70.5	118.5	2951
395	$AlCl_3$–KCl–LiCl	59.2–19.7–21.1	120.0	3252
396	AlI_3–HgI_2	65	120.0	1918
397	GaI_3–SbI_3	45	120.0	2289
398	GaI_3–SnI_4	70	120.0	2289

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References					
399	HgI_2–NH_4I	43	120.0	514					
400	$NaNO_3$–$TlNO_2$–$TlNO_3$	32–29–39	120.0	1687					
401	KCNS–KNO_2	66	120.0	1079					
402	KNO_3–$LiNO_3$–$NaNO_3$	44.5–37.5–18	120.0	1009					
403	KNO_3–$LiNO_3$–$NaNO_3$	44.9–37.3–17.8	120.0	1350	1351	1352			
404	KNO_3–$LiNO_3$–$NaNO_3$	53–30–17	120.0	965					
405	K_2WO_4–Li_2WO_4	56	120.0	2295					
406	$FeCl_3$–NaCl–$TeCl_4$	58.2–22.0–19.5	120.0	3021					
407	$AlCl_3$–$NaAlCl_4$–$TeCl_4$	NA	120.0	2834					
408	$AlCl_3$–$FeCl_3$–$MoCl_5$	NA	121.0	896					
409	$NaNO_3$–NH_4NO_3	20.5	121.0	1009					
410	AgCl–$HgCl_2$–HgI_2	25–26–50 APP	121.0	3113					
411	$AgNO_3$–$Cd(NO_3)_2$–KNO_3	48–10.5–41.5	121.0	3119					
412	$AlCl_3 \cdot BiCl_3$–$FeCl_3$–$NaAlCl_4$	21.3–5.3–73.4	122.0	2594					
413	$AlCl_3$–$HfCl_4$	98	122.0	1124					
414	CuCl–TlCl	60	122.0	715					
415	CsI–InI_3	18	122.0	1970					
416	$AlCl_3$–NaCl	59	123.0	27	34	44	45	78	80
				84	258	331	451	472	688
417	AlI_3–NaI	70	123.0	2523					
418	$KClO_4$–KNO_3–$LiNO_3$	1–56.5–42.5	123.0	2786					
419	$NaNO_2$–$Sr(NO_2)_2$–$TlNO_2$	13.4–14–72.6	124.0	1853					
420	KNO_3–$LiNO_3$	57.5	124.0	2295					
421	KNO_3–$LiNO_3$	59	124.0	917					
422	NaCNS–$RbNO_3$	33	124.0	1940					
423	$NaNO_3$–$TlNO_2$	25	124.5	1687					
424	$AlCl_3$–KCl–NaCl	50–15–35	125.0	840					
425	$AlCl_3$–$NbCl_5$	54.3	125.0	241	331	658			
426	$Ca(NO_3)_2$–$CO(NH_2)_2$–KNO_3	23.9–57.7–18.3	125.0	2117					
427	$LiNO_2$–$NaNO_2$–$NaNO_3$	55–12–33	126.0	916					
428	$LiNO_2$–$RbNO_3$	37.5	126.0	1012					
429	$CsC_2H_3O_2$–$CsNO_3$–$LiC_2H_3O_2$	71.5–4–24.5	126.0	1164					
430	$LiNO_3$–LiOH–$RbNO_3$–RbOH	NA	126.0	2825					
431	InI_3–SbI_3	47.5	127.0	2423					
432	$CsC_2H_3O_2$–$CsNO_3$–$LiC_2H_3O_2$	69.5–4–26.5	127.0	2902					
433	$AlCl_3$–KCl	67	128.0	34	43	44	45	78	80
				84	260	794	3241	3243	
434	$CsNO_2$–$LiNO_3$	32.5	128.0	1192					
435	$Ca(NO_3)_2$–$RbNO_3$	39	128.0	1998					
436	KNO_3–$LiNO_3$	62	128.0	2275					
437	$CsC_2H_3O_2$–$CsNO_3$–$RbC_2H_3O_2$	55–23–22	128.0	1278					
438	$AgNO_3$–$RbNO_3$	67.5	128.0	3121					
439	AsI_3–HgI_2	88.5	129.0	2452					
440	$AlCl_3$–$HfCl_4$–NaCl	46.5–4–49.5	130.0	1124					
441	$AlCl_3$–NaCl–WCl_6	48–50–2	130.0 ±3	1106					
442	AsI_3–GaI_3	6	130.0	2289					
443	$LiNO_2$–$LiNO_3$–$NaNO_3$	60–5–35	130.0	916					
444	$CsNO_3$–$LiNO_3$–$NaNO_3$	36–47.5–16.5	130.0	1211					
445	$LiNO_3$–$NaNO_3$–$RbNO_3$	28.5–20–51.5	130.0	1211					
446	AsI_3–InI_3	65	130.0	2682					
447	GeI_4–InI_3	84	131.0	2682					
448	NH_4NO_3–$Pb(NO_3)_2$	67 APP	131.0	3160					
449	KNO_3–NH_4Cl–NH_4NO_3	11–14–75	131.0	3187					
450	$AlCl_3$–KCl	71–65.5	131.5 ±17.5	688					
451	NH_4Cl–$NH_4H_2PO_4$–NH_4NO_3	14–6–80	131.3	3187					
452	$AgNO_3$–KNO_3	62	131.9	1943					
453	HgI_2–KI	61.5	132.0	1918					
454	AlI_3–HgI_2	42	132.0	1918					

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References		
455	$LiNO_2$–$RbNO_3$	65	132.0	1012		
456	$CsNO_3$–$LiNO_3$–$NaNO_3$	47–40.5–13.5	132.0	1211		
457	KNO_3–$LiNO_3$	58.8	132.0	1065		
458	$LiNO_3$–$TlNO_3$	30	132.0	1293		
459	$CsC_2H_3O_2$–$KC_2H_3O_2$	71.5	132.0	1213		
460	$LiNO_2$–$NaNO_3$	65	133.0	916		
461	$LiNO_3$–LiOH–$RbNO_3$–RbOH	NA	133.0	2825		
462	$LiNO_3$–LiOH–$RbNO_3$–RbOH	NA	133.0	2825		
463	$BiCl_3$–$NaFeCl_4$	17.5	134.0	2594		
464	$CsNO_2$–$LiNO_3$	47.5	134.0	1192		
465	$Cd(NO_3)_2$–KNO_3–$LiNO_3$	34.2–43–22.8	134.0	512		
466	KNO_3–$LiNO_3$	58.8	134.0	925		
467	$LiNO_3$–$NaNO_3$–$RbNO_3$	45–21–34	134.0	1211		
468	$CsNO_3$–CsOH–KOH	17–57–26	134.0	3054		
469	AlOCl–$NbCl_5$	42.9	135.0	2565		
470	$HgCl_2$–NH_4Cl	42	135.0	513	514	655
471	NaCl–$RbNO_3$–TlBr	17–79–4	135.0	956		
472	AlI_3–LiI	74 APP	135.0 APP	2284		
473	AlI_3–RbI	71	135.0	2523		
474	AlI_3–SbI_3	60	135.0	1918		
475	$Cd(NO_3)_2$–$NaNO_3$	47	135.0	904	1998	
476	$Mg(NO_3)_2$–$NaNO_3$	33	135.0	1998		
477	$CsC_2H_3O_2$–$LiNO_3$	85.1	135.0	1164		
478	$CsNO_3$–$LiC_2H_3O_2$–$LiNO_3$	39–11–50	135.0	1164		
479	$CsC_2H_3O_2$–$LiC_2H_3O_2$	76	135.0	952		
480	$AlCl_3$–$BiCl_3$–NaCl	48.7–10.1–41.2	136.0	2525		
481	$BiCl_3$–$FeCl_3$–$NaAlCl_4$	22.3–14–63.7	136.0	2594		
482	CuCl–KCl	67	136.0	38	104	715
483	$AlCl_3 \cdot POCl_3$–$ZrCl_4 \cdot 2POCl_3$	47	136.0	2359		
484	AgCl–HgI_2	52	136.0	1918		
485	$Sr(NO_2)_2$–$TlNO_3$–$TlNO_2$	26.5–11.0–62.5	136.0	2974		
486	$CsNO_3$–$LiC_2H_3O_2$–$LiNO_3$	5.5–51–43.5	136.0	2902		
487	AgCl–AgI–HgI_2	47–2–51 APP	136.0	3113		
488	$AgNO_3$–$RbNO_3$	40	136.0	3121		
489	$AlCl_3$–$NbCl_5$	41.6	136.5	1344		
490	$LiNO_3$–$TlNO_3$	31	136.5	925		
491	$AlCl_3$–$HfCl_4$–NaCl	46.5–2–51.5	137.0	1124		
492	$LiNO_2$–$NaNO_2$–$Sr(NO_2)_2$	55.1–36.7–8.2	137.0	2285		
493	$LiC_2H_3O_2$–$LiNO_3$–$NaNO_3$	49.5–42.5–8	137.0	1146		
494	InCl–$ZnCl_2$	55	137.0	2705		
495	AgCl–HgI_2	49 APP	137.0	3113		
496	$AlBr_3$–$BiBr_3$	34.4	137.3	1918		
497	$AlCl_3$–$BiCl_3$–NaCl	44–11–45	138.0	2525		
498	$AlCl_3$–$MoCl_5$–NaCl	47–3–50	138.0 ±2	912		
499	$AlCl_3$–InCl	70	138.0	2186	2392	
500	$AlCl_3$–$NbCl_5$	50	138.0	1182		
501	$CsC_2H_3O_2$–$CsNO_3$–$LiC_2H_3O_2$	16.5–8.5–75	138.0	1164		
502	$NaAlCl_4$–WCl_5	98	138.0	3015		
503	AlI_3–InI_2	27.5	138.0	2919		
504	$CsC_2H_3O_2$–$KC_2H_3O_2$	72	139.0	1322		
505	KNO_3–$NH_4H_2PO_4$–NH_4NO_3	NA	139.0	3187		
506	$AlCl_3$–$MoCl_5$–NaCl	48.5–3–48.5	140.0 ±2	912		
507	$AlCl_3$–NaCl–$NbCl_5$	48.8–48.4–2.8	140.0	331		
508	$AlCl_3$–NaCl–$TaCl_5$	46.7–51.2–2.1	140.0	121		
509	$BiCl_3$–$NaAlCl_3$	14	140.0	2594		
510	$HgBr_2$–NH_4Br	52	140.0	514		
511	AlI_3–NaI	71 APP	140.0 APP	2284		
512	GaI_3–KI	65	140.0	2243		

Locator number	System	Mol %	T, °C		References
513	AlI_3–SbI_3	34	140.0	1918	
514	GaI_3–HgI_2	70 APP	140.0 APP	2263	
515	InI_3–PbI_2	77.5	140.0	1345	
516	GaI_3–TlI	65	140.0	2243	
517	GaI_3–ZnI_2	38	140.0	2263	
518	$NaNO_2$–$TlNO_2$	18 APP	140.0	1148	
519	$CsNO_3$–KNO_3–$NaNO_3$	29–33–38	140.0	1214	
520	$Ca(NO_3)_2$–K_2CrO_4–KNO_3	34.21–0.02–65.77	140.0	546	
521	KNO_3–NaCNS	63.5	140.0	944	
522	KNO_3–NaCNS	34 APP	140.0 APP	1942	
523	$CsC_2H_3O_2$–$Zn(C_2H_3O_2)_2$	80	140.0	2957	
524	$NaAlCl_4$–$TeCl_4$	NA	140.0	2834	
525	CuCl–$FeCl_3$–$TeCl_4$	8–30–62	140.0	2918	
526	CuCl–NH_4Cl	62	140.0	3147	
527	GaI_3–SbI_3	27.5	140.1	2621	
528	$Ca(NO_3)_2$–$CO(NH_2)_2$	25.2	140.5	2109	
529	$NaC_2H_3O_2$–$RbC_2H_3O_2$–$RbNO_3$	31–60–9	141.0	1030	
530	$AgNO_3$–$RbNO_3$	5	141.0	3121	
531	$AlCl_3$–$FeCl_3$–NaCl	49–2–49	142.0	490	
532	$FeCl_3$–$NaAlCl_4$	2	142.0	2594	
533	$AlCl_3 \cdot POCl_3$–$HfCl_4 \cdot 2POCl_3$	49	142.0	2359	
534	$Sr(NO_2)_2$–$TlNO_2$	12.7	142.0	910	
535	$Ca(NO_3)_2$–KNO_3	34.2	142.0	546	
536	$CsC_2H_3O_2$–$CsNO_3$	75	142.0	1213	
537	$CsNO_2$–$NaNO_2$	50	142.0	2923	
538	$NaC_2H_3O_2$–$Zn(C_2H_3O_2)_2$	43	142.0	3109	
539	CsN_3–$Zn(N_3)_2$	NA	142.0	3072	
540	AlI_3–InI	65	142.0	2919	
541	AgI–Ag_2SO_4	68	142.0	3117	
542	CuCl–NH_4Cl	41	143.0	61	224
543	$Ca(NO_3)_2$–KNO_3–$Sr(NO_3)_2$	27.1–69.4–3.5	143.0	1237	
544	$NaAlCl_4$–$ReCl_5$	95 APP	143.0	3015	
545	$NaAlCl_4$–$NbOCl_3$	98	143.6	2086	
546	$FeCl_3$–NaCl–WCl_6	48–48–4 APP	144.0 ±2	1106	
547	$KC_2H_3O_2$–$NaC_2H_3O_2$–$RbC_2H_3O_2$	8–33–59	144.0	996	
548	$AlCl_3$–$FeCl_3$–NaCl	48–3–49	145.0	490	
549	InCl–$ZnCl_2$	57.6	145.0	964	
550	$HgCl_2$–HgI_2	55	145.0	1405	
551	$Ca(NO_3)_2$–KNO_3	34.2	145.0	1998	
552	$Ca(NO_3)_2$–KNO_3	36	145.0	1789	
553	$LiC_2H_3O_2$–$LiNO_3$	51	145.0	1146	
554	$NaC_2H_3O_2$–$NaNO_3$–$RbNO_3$	10–38.5–51.5	145.0	1030	
555	AlI_3–CsI–NaI	57.0–21.0–22.0	145.0	2715	
556	NaCl·$AlCl_3$–NaI	95 APP	145.0	3237	
557	$CsNO_3$–$LiNO_3$–$RbNO_3$	7–64–29	145.0	3079	
558	$AlCl_3$–NaCl–$TeCl_4$	44.5–46.5–9	145.0	2834	
559	$Ca(NO_3)_2$–KNO_3	34 APP	145.0	3174	
560	GaI_3–SbI_3	55.1	145.3	2621	
561	$Ca(NO_3)_2$–KNO_3	32.4	146.0	1493	
562	$LiNO_3$–LiOH–$RbNO_3$–RbOH	NA	146.0	2825	
563	$LiNO_2$–$LiNO_3$	70	147.0	1191	
564	$CsC_2H_3O_2$–$LiNO_3$	21.8	147.0	1164	
565	$BiCl_3$–$GaCl_3$	65	147.0	2964	
566	$CsNO_3$–$LiNO_3$–$RbNO_3$	23–40–37	147.0	3079	
567	HgI_2–SbI_3	22.5	147.5	2452	
568	$FeCl_3$–NaCl–$NbCl_5$	48.–50.5–1.5	148.0	446	
569	$FeCl_3$–NaCl–$NbCl_5$	52.3–46.0–1.7	148.0	446	
570	$AlCl_3$–CsCl	74.6	148.0	240	

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
571	HgI_2–SbI_3	24.3	148.0	1448
572	$LiNO_3$–$RbNO_2$	67.5	148.0	1012
573	$NaNO_2$–$TlNO_3$	60	148.0	1687
574	KNO_3–$NaNO_3$–$RbNO_3$	25.5–37–37.5	148.0	1212
575	$LiNO_3$–$RbNO_3$	30	148.0	1211
576	$CsNO_3$–NaCNS	61	148.0	1940
577	$LiC_2H_3O_2$–$NaC_2H_3O_2$–$NaNO_3$	51–35.3–13.5	148.0	1146
578	CsN_3–$Zn(N_3)_2$	NA	148.0	3072
579	$HgCl_2$–HgI_2	52	148.0	3113
580	$AgNO_3$–$Cd(NO_3)_2$	62 APP	148.0	3119
581	$AlCl_3$–$FeCl_3$–NaCl	46–3–51	149.0	490
582	$AlBr_3$–CsBr	79	149.0	2265
583	HgI_2–TlI	60 APP	149.0	1844
584	$LiNO_2$–$NaNO_2$	63	149.0	1201
585	$LiNO_3$–$NaNO_2$	35	149.0	916
586	$AlCl_3$–NaCl	50	150.0	27 34 44 45 78 80 84 258 331 451 472 688
587	CuCl–KCl	66	150.0	38 104 715
588	CuCl–RbCl	68	150.0	714
589	GaI_3–NaI	67	150.0	2243
590	$CsNO_3$–$NaNO_3$	48	150.0 APP	3239
591	$CsNO_3$–$NaNO_3$–$RbNO_3$	16.5–40–43.5	150.0	1214
592	$CsI{\cdot}2AlI_3$–$NaI{\cdot}AlI_3$	46 APP	150.0	2715
593	$BiCl_3$–CuCl–$FeCl_3$	36–17–43	150.0	2918
594	$BiCl_3$–TlCl	67.5	150.0	3133
595	$AlCl_3$–$TaCl_5$	49.8	150.6	2187
596	$FeCl_3$–NaCl	54.7	151.0	263
597	$LiNO_2$–$NaNO_2$	62.5	151.0	916
598	$NaNO_2$–$TlNO_3$	25	151.0	1687
599	CsF–HF	54.7	151.5	566
600	$NaBiCl_4$–$NaFeCl_4$	20	152.0	2594
601	$Cd(NO_3)_2$–$RbNO_3$	23	152.0	1998
602	$NaC_2H_3O_2$–$RbC_2H_3O_2$–$RbNO_3$	15.5–35–49.5	152.0	1030
603	AgI–InI_3	10	153.0	1970
604	$AgIO_3$–$AgNO_3$	26.1 APP	153.0 APP	1094
605	$Ba(NO_2)_2$–$RbNO_2$–$TlNO_2$	20–33–47	153.0	2725
606	$AlCl_3$–$BiCl_3$–NaCl	50–0–50 APP	154.0	2525
607	$SnCl_2$–$TaCl_5$	69.5	154.0	914
608	$NaNO_3$–Na_2SO_4–$RbNO_3$	40.4–1.0–58.6	154.0	2529
609	$Ca(NO_2)_2$–$NaNO_2$–$NaNO_3$	26.6–32.9–40.5	154.0	915
610	$LiNO_3$–$RbNO_3$	32	154.0	1082
611	$CsNO_3$–LiBr–$LiNO_3$	41–9–50	154.0	2615
612	Ga–$GaBr_3$	36	154.0	3018
613	CdI_2–CsI–KI	46–19–35	155.0	1794
614	$BiCl_3$–$SeCl_4$	50	155.0	2619
615	NH_4NO_3–$(NH_4)_4P_2O_7$	95.4	155.0	3058
616	AgBr–$AgNO_3$	25	155.0	3168
617	$LiNO_3$–$NaNO_2$	62.5	156.0	916
618	$BiCl_3$–PCl_5	NA	156.0	2930
619	KNO_3–NH_4NO_3	21.3	156.7	1009
620	KNO_3–NH_4NO_3	11.3	156.7	3187
621	$FeCl_3$–NaCl–$WOCl_4$	50–50–0 APP	157.0	2467
622	$NbCl_5$–WCl_6	56.7	157.0	1455
623	$KC_2H_3O_2$–$LiC_2H_3O_2$–$NaC_2H_3O_2$	41–32–27	157.0	1145
624	$KC_2H_3O_2$–$LiC_2H_3O_2$–$NaC_2H_3O_2$	9–57–34	157.0	1145
625	$FeCl_3$–NaCl	54	158.0	121 372
626	KCl–$NbCl_5$–$ZrCl_4$	17–44–39	158.0	874
627	$TaCl_5$–WCl_6	49.9	158.0	1455

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References
628	$Ca(NO_2)_2$–$TlNO_2$	8.1	158.0	910	
629	$Ba(NO_3)_2$–$Ca(NO_3)_2$–KNO_3	33.3–17.5–49.1	158.0	1237	
630	$KC_2H_3O_2$–$LiC_2H_3O_2$–$NaC_2H_3O_2$	44–31–25	158.0	1145	
631	$GaCl_3$–PCl_5	100 APP	158.0	2648	
632	$AlCl_3$–WCl_5	70	158.0	3015	
633	$KBiCl_4$–$LiBiCl_4$	35	158.0	3042	
634	$RbC_2H_3O_2$–$Zn(C_2H_3O_2)_2$	40	159.0	3109	
635	$AlCl_3$–$BiCl_3$	35	160.0	2525	
636	$AlCl_3$–$TlCl$	70.5	160.0	1157	
637	GaI_3–SnI_2	82	160.0	2263	
638	$CsNO_3$–$TlNO_2$	20	160.0	1242	
639	$LiC_2H_3O_2$–$NaC_2H_3O_2$	57	160.0	1146	
640	AlI_3–CsI–NaI	48.0–21.0–31.0	160.0	2715	
641	$CsNO_3$–KNO_3–KOH	22–53–25	160.0	3054	
642	$(NH_4)_4P_2O_7$–$(NH_4)H_2PO_4$	27.5	160.0	3058	
643	$AgCl$–$AgNO_3$	25	160.0	3168	
644	$AgCl$–$AgNO_3$	25	160.0	3168	
645	KNO_2–KOH	58.4	160.0	3194	
646	$LiNO_3$–$LiOH$–$NaNO_3$	53–33–14	160.0	3211	
647	$InBr_3$–$TlBr$	NA	161.0	3076	
648	$Ba(NO_2)_2$–KNO_2–$NaNO_2$	21.2–42.4–36.4	162.0	2123	
649	$CsNO_2$–$TlNO_2$	25	162.0	1148	
650	$Ca(NO_3)_2$–KNO_3	29.5	162.0	1237	
651	$NaNO_3$–$TlNO_3$	20.5	162.0	235	
652	$NaNO_3$–$TlNO_3$	22	162.0	1293	
653	$KC_2H_3O_2$–$LiC_2H_3O_2$–$NaC_2H_3O_2$	38–32–30	162.0	1218	
654	$LiClO_4$–$LiNO_3$–$NaClO_4$	60–26–14	162.0	2710	
655	$KBiCl_4$–$LiBiCl_4$	60	162.0	3042	
656	$LiNO_3$–$NaClO_4$–$NaNO_3$	43–19.5–37.5	162.0	2774	
657	$POCl_3$–$ZrCl_4$	59.75	162.5	674	2555
658	$Cd(NO_3)_2$–$CsNO_3$	39.4	162.5	2181	
659	CoI_2–InI_2	15	163.0	2994	
660	$AgNO_3$–$CsNO_3$	67.5	163.0	3120	
661	$HgBr_2$–$TlBr$	60 APP	164.0	1844	
662	AlI_3–CsI	85	164.0	2523	
663	$NaNO_3$–$RbNO_3$	43	164.0	1211	
664	$NaNO_3$–$TlNO_3$	23.4	164.2	1943	
665	$AlCl_3$–$POCl_3$	41 APP	164.4	1159	
666	$AlCl_3$–$POCl_3$	40.4	164.5	1182	
667	$AlCl_3$–$POCl_3$	41 APP	164.5	2359	
668	$AlCl_3$–$ZrCl_4$	85.7	165.0	1429	
669	BiI_3–InI_3	8	165.0	1878	
670	$Ba(NO_2)_2$–$TlNO_2$	8.1	165.0	910	
671	$Ca(NO_2)_2$–$Ca(NO_3)_2$–$NaNO_3$	35.1–12.1–52.7	165.0	915	
672	$CsNO_3$–$RbC_2H_3O_2$–$RbNO_3$	30–44–26	165.0	1278	
673	$NaCNS$–$TlNO_3$	15 APP	165.0 APP	1942	
674	$CsI \cdot AlI_3$–$NaI \cdot AlI_3$	43 APP	165.0	2715	
675	$Ca(NO_3)_2$–$LiNO_3$–$NaNO_3$	0.18–99.45–0.37	165.0	2943	
676	$KClO_4$–$LiClO_4$–$LiNO_3$	5–52.5–42.5	165.0	2786	
677	$BiCl_3$–$TeCl_4$	49.3	165.5	3229	
678	$NaCl$–Na_2SO_4–$KCNS$	3–0.3–96.7	166.0	246	
679	$NaNO_3$–Na_2SO_4–$RbNO_3$	28.1–0.5–71.3	166.0	2529	
680	$Cd(NO_3)_2$–$CsNO_3$	29.4	166.0	2181	
681	$InCl$–$ZnCl_2$	85	166.0	2705	
682	$Cd(C_2H_3O_2)_2$–$CsC_2H_3O_2$	15	167.0	3110	
683	KNO_2–KOH	43	167.0	3194	
684	KCl–$KCNS$–K_2SO_4	3.4–96.4–0.2	168.0	246	
685	$HgBr_2$–KBr	57.5	168.0	343	

Locator number	System	Mol %	T, °C	References	
686	KNO_2–$TlNO_2$	25	168.0	1148	
687	$Ba(NO_2)_2$–$Ba(NO_3)_2$–$LiNO_2$	15.2–4.2–80.6	168.0	1161	
688	$Cd(NO_3)_2$–KNO_3	46	168.0	1998	
689	$Cd(NO_3)_2$–$RbNO_3$	42	168.0	1998	
690	$Ca(NO_3)_2$–K_2CrO_4–KNO_3	28.62–0.02–71.36	168.0	546	
691	$(NH_4)_4P_2O_7$–$(NH_4)_2HPO_4$	41±5	168.0 ±2	2891	
692	KCNS–KI	93 APP	168.0	3181	
693	$AgNO_3$–$CsNO_3$	82.5	168.5	3120	
694	KCNS–NaCl	97	169.0	246	
695	$KC_2H_3O_2$–$Zn(C_2H_3O_2)_2$	30	169.0	3109	
696	$AlCl_3$–KCl–$NbCl_5$–$TaCl_5$	18–50–15.5–16.5	170.0	332	
697	$BiCl_3$–KCl	82.5 ±2.5	170.0	1918	
698	$AlCl_3$–$WOCl_4$	75.7	170.0 ±2.	2467	
699	$AlCl_3$–$ZrCl_4$	97.5 APP	170.0	1189	
700	CuI–InI_3	5	170.0	901	
701	KOH–NaOH	50	170.0	2037	2178
702	KOH–NaOH	50.6	170.0	829	
703	$Ca(NO_2)_2$–$RbNO_3$	19.7	170.0	1129	
704	$CsNO_3$–$LiNO_3$	40	170.0	1192	
705	$LiNO_3$–$RbNO_3$	61.5	170.0	1211	
706	$MoCl_5$–$SeCl_4$	58	170.0	2619	
707	$(NH_4)_2P_2O_7$–$(NH_4)_2HPO_4$	36.1	170.0	3058	
708	KCl–$TiCl_3$–$ZrCl_4$	NA	170.0	2837	
709	SbI_3–Sb_2O_3	NA	170.0	2863	
710	KCl–KCNS	NA	170.2	3181	
711	KBr–KCNS	3 APP	170.2	3181	
712	$BiCl_3$–$FeCl_3$	63	171.0	2594	
713	$SnCl_2$–$ZnCl_2$	56	171.0	61	
714	$SnCl_2$–$ZnCl_2$	64	171.0	3138	
715	$AgNO_3$–$LiNO_3$	75	171.5	3121	
716	$BiCl_3$–$FeCl_3$	77	171.5	3138	
717	CuCl–$SnCl_2$	21.8	172.0	1918	
718	NH_4Cl–$SnCl_2$	19	172.0	834	
719	$TaCl_5$–$WOCl_4$	50.2	172.0	2054	
720	AgCl–$AgNO_3$–$Ca(NO_3)_2$	18.5–78.3–3.1	172.0	198	
721	$NaNO_3$–$RbNO_3$	28	172.0	1211	
722	KN_3–$Zn(N_3)_2$	NA	172.0	3072	
723	CuCl–$SnCl_2$	18	172.0	3138	
724	$AlCl_3$–$BiCl_3$	68	173.0	2525	
725	$HfCl_4$–$POCl_3$	40	173.0	2555	
726	NH_4Cl–$SnCl_2$	42	173.0	834	
727	$BiBr_3$–$BiCl_3$	49.5	173.0	2014	
728	KCl–KCNS	3.4	173.0	246	291
729	CdI_2–CsBr	56.9	173.0	1010	
730	CdI_2–CsI–NaI	37.6–50.4–12	173.0	1795	
731	$AgNO_3$–$LiNO_3$	76	173.0	925	
732	$CsC_2H_3O_2$–$RbC_2H_3O_2$	72 APP	173.0	1013	
733	CdI_2–KI–PbI_2	42.9–49.6–7.5	173.0	3023	
734	CsCl–$SnCl_2$	17	174.0	834	
735	$NbCl_5$–$WOCl_4$	56.3	174.0	2054	
736	LiCl–$LiNO_3$–$NaNO_3$	5–80–15	174.0	2393	
737	$CsNO_3$–$LiNO_3$	43	174.0	1082	
738	$MoCl_5$–$SeCl_4$	64	174.0	2686	
739	$CsNO_3$–KNO_3–KOH	26.5–20.5–53	174.0	3054	
740	$HgBr_2$–TlBr	48 APP	175.0	1844	
741	GaI_3–TeI_4	27	175.0	2289	
742	$Cd(NO_3)_2$–KNO_3	25	175.0	1998	
743	$LiNO_3$–NaCNS	56 APP	175.0 APP	1942	

Locator number	System	Mol %	T, °C		References			
744	$NaCNS-RbNO_3$	62	175.0	1940				
745	$CsI·AlI_3-KI·AlI_3$	41 APP	175.0	2715				
746	$POCl_3-ZrCl_4$	43.5	175.3	674	2555			
747	$KCl-SnCl_2$	52	176.0	834				
748	$AlCl_3-ZrCl_4$	92.5	176.0	1876				
749	$AgCl-AgNO_3$	18.5	176.0	61	62	175	341	376
750	$HgCl_2-TlNO_3$	12.8	176.0	476				
751	$RbNO_3-TlNO_2$	10	176.0	2101				
752	$LiC_2H_3O_2-RbC_2H_3O_2$	26	176.0	952				
753	$Pb(NO_3)_2-TlNO_3$	12.1	176.8	1943				
754	$CsNO_3-NaNO_3$	47	177.0	1213				
755	$FeCl_3-PbCl_2$	63	177.0	3138				
756	$AlCl_3-FeCl_2$	87	178.0	1973				
757	$SnCl_2-TlCl$	70.5	178.0	512				
758	$Ca(NO_3)_2-LiNO_2-LiNO_3$	10.5–71.8–17.7	178.0	1055				
759	$Ca(NO_3)_2-RbNO_2$	17.6	178.0	1129				
760	$KNO_3-Mg(NO_3)_2$	56	178.0	1998				
761	$CsNO_3-LiBr-LiNO_3$	26–65–9	178.0	2615				
762	$CsNO_3-LiBr-LiNO_3$	36–56.5–7.5	178.0	2615				
763	$LiBiCl_4-NaBiCl_4$	85	178.0	2962				
764	$Te-TeI_4$	66	178.0	2851				
765	$TaCl_5-WOCl_4$	55	178.5	1993				
766	$NaNO_3-RbNO_3$	44	178.5	1082				
767	$FeCl_3-PbCl_2$	50	178.6	3138				
768	$NaC_2H_3O_2-RbC_2H_3O_2$	38	179.0	1030				
769	$Cd(C_2H_3O_2)_2-RbC_2H_3O_2$	14	179.0	3110				
770	$LiNO_3-RbNO_3$	66	179.5	1082				
771	$NaF-SnF_2$	20	180.0	1306				
772	$KCl-NaCl-TaCl_5$	20–30–50	180.0	840				
773	$AlCl_3-KCl-NbCl_5$	4–6–90	180.0	332				
774	$AlCl_3-KCl-TaCl_5$	21.3–50–28.7	180.0	332				
775	$FeCl_3-KCl-ZrCl_4$	30–35–35	180.0	794				
776	$FeCl_3-KCl-ZrCl_4$	59–34–7	180.0	794				
777	$KCl-NbCl_5-TaCl_5$	15.5–69–15.5	180.0	332				
778	$KCl-SnCl_2$	38	180.0	502	667			
779	$BiCl_3-TaCl_5$	73	180.0	2328				
780	$NH_4Cl-ZnCl_2$	48.5	180.0	964	965			
781	$SnCl_2-ZnCl_2$	58.2	180.0	964				
782	$LiCl-LiNO_3-NaNO_3$	2.5–52.5–45	180.0	2393				
783	$CdI_2-KI-NaI$	47.1–42.6–10.3	180.0	2148				
784	$AgI-GaI_3$	24	180.0	2243				
785	BiI_3-GaI_3	31	180.0	2289				
786	CdI_2-GaI_3	20 APP	180.0 APP	2263				
787	GaI_3-PbI_2	85	180.0	2263				
788	$RbNO_2-TlNO_2$	75	180.0	1148				
789	$Ca(NO_2)_2-NaNO_3$	36.8	180.0	915				
790	$MoCl_5-PCl_5$	62	180.0	2686				
791	$LiNO_3-NaClO_4-NaNO_3$	43–19.5–37.5	180.0	2710				
792	$FeCl_3-WCl_5$	70	180.0	3015				
793	$AlCl_3-BaCl_2$	81.7	180.0	3034				
794	$LiClO_4-LiNO_3-NaClO_4$	60–26–14	180.0	2774				
795	$LiI-LiOH$	55	180.0	3202				
796	$NaNO_3-RbNO_3$	53	180.5	1077				
797	$Ba(NO_2)_2-NaNO_2$	55.6	181.0	2123				
798	KNO_3-TlNO_3	28	181.0	1293				
799	$RbC_2H_3O_2-RbNO_3$	35.5	181.0	1030				
800	$KC_2H_3O_2-LiC_2H_3O_2$	64	181.0	1145				
801	$KNO_3-TlBr-TlNO_3$	25.5–4.5–70	181.0	3195				

Locator number	System	Mol %	T, °C	References
802	$AlCl_3$–KCl–$ZrCl_4$	20–40–60	182.0	794
803	$HgCl_2$–KCl	68	182.0	513
804	$LiClO_4$–NH_4ClO_4	69.5	182.0	1116
805	CuBr–KBr	28.1	182.0	641
806	$Ba(NO_2)_2$–$LiNO_2$	20.5	182.0	1161
807	InI_3–ZnI_2	75	182.0	2616
808	NaCl–$SnCl_2$	32	183.0	667
809	$HgCl_2$–TlCl	64	183.0	711
810	NH_4Cl–$SnCl_2$	20	183.0	953
811	$CsNO_3$–$LiNO_3$	55	183.0	1192
812	$NaC_2H_3O_2$–NaCNS–$NaNO_3$	32.5–30.5–37.0	183.0	3230
813	$HfCl_4$–$POCl_3$	56	183.7	2555
814	$CeCl_3$–NaCl–$SnCl_2$	2–22–76	184.0	828
815	NaCl–$NbCl_5$–$ZrCl_4$	18.8–54.4–26.7	184.0	779
816	$CsNO_3$–NaCNS	31	184.0	1940
817	AgCl–$PbCl_2$–TlCl	50–8–42	184.0	2651
818	AgCl–$BiCl_3$	45	184.0	2964
819	NH_4Cl–$NH_4H_2PO_4$	12.2	184.0	3187
820	$SnCl_2$–TlCl	82	185.0	711
821	AlI_3–CsI	64	185.0	2523
822	$Ca(NO_2)_2$–KNO_2	25	185.0	1897
823	CdI_2–KI	52.5	185.0	3140
824	NaCl–$SnCl_2$	28	186.0	828
825	$NbCl_5$–$ZrCl_4$	65.7	186.0	779
826	NH_4Cl–$SnCl_2$	40	186.0	953
827	$KC_2H_3O_2$–$RbNO_3$	25	186.0	1225
828	$Ba(NO_2)_2$–$LiNO_2$	21.2	187.0	2116
829	$LiNO_3$–$NaNO_3$	56	187.0	3003
830	CsCl–$SnCl_2$	17	188.0	286
831	$AlCl_3$–InCl	80	188.0	2186 2392
832	FeI_2–GaI_3	25	188.0	2263
833	GaI_3–NiI_2	71	188.0	2263
834	$Ca(NO_3)_2$–$CsNO_3$	42	188.0	1998
835	CsOH–KOH	55	188.0	3054
836	KNO_3–$NH_4H_2PO_4$	33	188.0	3187
837	$AlCl_3$–$HfCl_4$	93.95	189.0	1881
838	$AlCl_3$–$ZrCl_4$	89.84	189.0	1881
839	$HgCl_2$–TlCl	66 APP	189.0	1844
840	KCl–NaCl–$NbCl_5$	22–28–50	190.0	840
841	$BiCl_3$–CuCl	56	190.0	1918
842	$BiCl_3$–$NbCl_5$	54	190.0	2328
843	AgCl–TlBr	60	190.0	1021
844	$AlBr_3$–KBr	48	190.0	2265
845	CdI_2–InI_3	62.5	190.0	1345
846	Na_2SO_4–$Na_2S_2O_7$	8	190.0	2009
847	$Cd(NO_3)_2$–$LiNO_3$	35	190.0	1998
848	$KC_2H_3O_2$–KNO_3–$LiC_2H_3O_2$	60–11–29	190.0	1215
849	$MoCl_5$–$SeCl_4$	45	190.0	2686
850	$BiCl_3$–$HgCl_2$	62	190.0	2964
851	$LiC_2H_3O_2$–$NaC_2H_3O_2$	26	190.0	3100
852	$GaBr_3$–TlBr	45	190.0	2809
853	KF–SnF_2	19	191.0	1306
854	$NbCl_5$–PCl_5	67.	191.0	2328
855	$SnCl_2$–TlCl	79	191.0	2061
856	RbCl–$SnCl_2$	17	192.0	286
857	$HgCl_2$–$TlNO_3$	37	192.0	476
858	HgI_2–$TlNO_3$	5	192.0	1404
859	$Ba(NO_3)_2$–$LiNO_3$–$NaNO_3$	1.3–53.2–45.5	192.0	512

Locator number	System	Mol %	T, °C		References	
860	$LiBiCl_4$–$NaBiCl_4$	50	192.0	2962		
861	$BiCl_3$–LiCl	95	192.0	2980		
862	$HgCl_2$–NH_4Cl	39	193.0	513	514	655
863	TlCl–$ZnCl_2$	41.8	193.0	964		
864	$LiNO_2$–$LiNO_3$	75	193.0	1161		
865	$LiNO_3$–$NaNO_3$	54	193.0	925		
866	TlBr–$TlNO_3$	11.5	193.0	3195		
867	$AlCl_3$–CsCl–$TaCl_5$	5–6–89	194.0	240		
868	$CsNO_3$–$NaNO_3$	50	194.0	2923		
869	AlI_3–InI	48	194.0 APP	2919		
870	$AlCl_3$–$HfCl_4$–KCl	22.5–46.2–31.3	195.0	1124		
871	$AlCl_3$–KCl–$NbCl_5$	15–52–33	195.0	332		
872	KNO_3–$Mg(NO_3)_2$	81	195.0	1998		
873	$CsI·AlI_3$–$RbI·AlI_3$	25 APP	195.0	2715		
874	$BeCl_2$–KCl–NaCl	46–13–41	195.0	2978		
875	$SnBr_2$–SnS	95 APP	195.0	2997		
876	$SnCl_2$–TlCl	81.0	195.2	2748		
877	$BiCl_3$–NaCl	80	196.0	2525		
878	NaCl–$NbCl_5$–$ZrCl_4$	35.6–49.2–15.2	196.0	779		
879	CdI_2–InI_3	1.5	196.0	1345		
880	$LiNO_3$–$NaNO_3$	55	196.0	916		
881	$Cd(C_2H_3O_2)_2$–$KC_2H_3O_2$	46	196.0	3110		
882	AgI–RbI	75	196.0	3116		
883	$AlCl_3$–$FeCl_2$–$FeCl_3$	59–33–8	197.0	1973		
884	$HgCl_2$–$TlNO_3$	58.5	197.0	476		
885	$LiC_2H_3O_2$–$NaNO_3$	85	197.0	1146		
886	$KC_2H_3O_2$–$LiC_2H_3O_2$	62.4	197.0	1218		
887	AgI–InI	70	197.0	2919		
888	RbCl–$SnCl_2$	18	198.0	834		
889	$FeCl_3$–$NbCl_5$	20.6	198.0	658		
890	$HgCl_2$–NH_4Cl	61.3	198.0	513	514	655
891	AlI_3–RbI	48	198.0	2523		
892	$Ca(NO_2)_2$–KNO_2	40.3	198.0	1897		
893	$RbC_2H_3O_2$–$RbNO_3$	81.5	198.0	1030		
894	KCl–$SnCl_2$	20	200.0	834		
895	RbCl–$SnCl_2$	44	200.0	834		
896	$AlCl_3$–InCl	30	200.0	2186	2392	
897	$FeCl_3$–$TaCl_5$	24.3	200.0	658		
898	NH_4Cl–$SnCl_2$	56	200.0	834		
899	GaI_3–KI	47	200.0	2243		
900	CdI_2–CsI	55.6	200.0	1010		
901	CuI–GaI_3	12.5	200.0	2243		
902	GaI_3–MgI_2	17	200.0	2263		
903	K_2CrO_4–KOH–LiOH	6–69–25	200.0	942		
904	$Ca(NO_2)_2$–$NaNO_2$	32.4	200.0	915		
905	$LiNO_2$–$LiNO_3$	70 APP	200.0 APP	1012		
906	$RbNO_3$–$Sr(NO_3)_2$	80	200.0	1998		
907	$RbNO_3$–$TlNO_3$	23 APP	200.0 APP	1293		
908	$NaC_2H_3O_2$–$NaNO_3$	76.5	200.0	1146		
909	$InBr_3$–NaBr	47	200.0	2625		
910	$CdCl_2$–InCl	10	200.0	2705		
911	$NaC_2H_3O_2$–$Zn(C_2H_3O_2)_2$	75	200.0	2957		
912	InCl–KCl	98 APP	200.0 APP	2767		
913	$BiBr_3$–$TeBr_4$	80	200.0	2841		
914	$InBr_3$–$SnBr_2$	27	200.0	2888		
915	AgI–TlI–Tl_2SO_4	1.5% SO4, 68% AG	200.0	3117		
916	$AlCl_3$–KCl–$NbCl_5$	27.5–50–22.5	201.0	332		
917	KCl–$SnCl_2$	17	201.0	502	667	

Locator number	System	Mol %	T, °C	References
918	$TeCl_4$–TlCl	89	201.0	2707
919	Tl_2CO_3–$TlNO_3$	14	201.0	3217
920	$FeCl_3$–KCl	53	202.0	794
921	AgBr–TlCl	61	202.0	1021
922	$Ca(ClO_4)_2$–$LiClO_4$–$NaClO_4$	11.2–66.8–22	202.0	913
923	$Ca(NO_3)_2$–$LiNO_2$	8.1	202.0	1055
924	$BiCl_3$–LiCl	65	202.0	2980
925	$FeCl_3$–KCl–UCl_4	NA	202.0	2985
926	$TeCl_4$–TeI_4	65	202.0 SER SOLID SOL	2764
927	$TeBr_4$–$TeCl_4$–TeI_4	0–65–35	202.0 SER SOLID SOL	2764
928	$AgNO_3$–$Ca(NO_3)_2$	89	202.0	3118
929	$NbCl_5$–$NbOCl_3$	99.4	202.4	2288
930	$AgNO_3$–$Pb(NO_3)_2$	96	202.7	1943
931	RbCl–$SnCl_2$	44.5	203.0	286
932	$RbNO_2$–$TlNO_3$	15	203.0	2101
933	KNO_3–$RbC_2H_3O_2$	20	203.0	1225
934	$GaCl_3$–KCl–$MgCl_2$	45–53–2	203.0	2613
935	KN_3–$Zn(N_3)_2$	NA	203.0	3072
936	KNO_3–$LiC_2H_3O_2$–$NaNO_3$	8–67–25	203.0	2901
937	$SnCl_2$–TlCl	66	204.0	2061
938	$LiClO_4$–$NaClO_4$	71.5	204.0	1116 2503
939	$Ca(NO_3)_2$–$NaNO_2$	31.1	204.0	1897
940	$LiNO_3$–$NaClO_4$	67	204.0	2710
941	$LiNO_3$–$NaNO_3$	67	204.0	2774
942	$NaNbOCl_4$–$NbCl_5$	0 APP	204.5 APP	2086
943	$KAlCl_4$–KCl–K_2NbOCl_5	75.4–11.5–13.1	205.0	1048
944	CuCl–$NbCl_5$	23.3 LT	205.0	889
945	$HgCl_2$–TlCl	32	205.0	711
946	$BiBr_3$–$PbBr_2$	76.5	205.0	1918
947	$Ca(NO_2)_2$–$LiNO_2$	9.6	205.0	1897
948	$Ca(NO_2)_2$–$RbNO_2$	21.2	205.0	1897
949	$RbNO_2$–$TlNO_3$	50	205.0	2101
950	$BaCl_2$–NaCl–$ZnCl_2$	6.5–32.5–61	205.0	2928
951	$FeCl_3$–KCl–UCl_4	NA	205.0	2985
952	K_2NbOCl_5–$TaCl_5$	57.5 APP	205.0	3025
953	AgI–TlI	69	205.0	3117
954	$BiBr_3$–$PbBr_2$	80	205.3	3138
955	$BiCl_3$–$WOCl_4$	40.3	206.0	2054
956	$AgNO_3$–$SrCl_2$	99.5	206.0	198
957	NaBr–NaI–NaOH	11–14–75	206.0	512
958	KNO_3–$RbC_2H_3O_2$	70	206.0	1225
959	$Ba(NO_2)_2$–KNO_2–$RbNO_2$	48–47–5	206.0	2725
960	$BiCl_3$–LiCl	45	206.0	2980
961	$AlCl_3$–KCl–$TaCl_5$	29–50–21	207.0	840
962	$KClO_4$–$LiClO_4$	24	207.0	495
963	$LiClO_4$–$NaClO_4$	71.5	207.0	913
964	AlI_3–KI	47.5	207.0	2523
965	$AgNO_3$–$Ba(NO_3)_2$	98	207.0	3118
966	$AgNO_3$–$Ba(NO_3)_2$	99.1	207.9	1942
967	$AlCl_3$–$FeCl_2$	58	208.0	1973
968	$SnCl_2$–$TaCl_5$	11.1	208.0	914
969	$LiClO_4$–$NaClO_4$	72.5	208.0	495
970	GaI_3–TlI	45	208.0	2243
971	$Ca(NO_2)_2$–$LiNO_3$	31.1	208.0	1055
972	$Ca(NO_3)_2$–$LiNO_2$–$LiNO_3$	30.7–61.4–7.8	208.0	1055
973	$CoCl_2$–NaCl–$TeCl_4$	2–3–95	208.0	3017
974	KNO_3–$NaNO_3$–$Sr(NO_3)_2$	52.3–42.9–4.7	208.4	1872
975	$AgNO_3$–$Sr(NO_3)_2$	99.5	208.7	1942

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References	
976	$FeCl_3$–$WOCl_4$	1 APP	209.0 ±2	2467		
977	$HgCl_2$–TlCl	40 APP	209.0	1844		
978	PCl_5–$TaCl_5$	30	209.0	2328		
979	$BeCl_2$–NaCl	55	210.0	218	287	314
980	$GaCl_3$–KCl	49	210.0	1016		
981	KCl–$PbCl_2$–$ZnCl_2$	42–6–52	210.0	682		
982	AgCl–TlCl	60	210.0	259	269	388
983	$AlCl_3$–$InCl_2$	3 APP	210.0 APP	2392		
984	$CdCl_2$–CdI_2–TlI	18.6–32.9–48.5	210.0	790		
985	GaI_3–NaI	47	210.0	2243		
986	LiOH–NaOH	30	210.0	1029	2037	
987	$AgNO_3$–$Pb(NO_3)_2$	93 APP	210.0	2115		
988	$RbNO_3$–$Sr(NO_3)_2$	84.8	210.0	955		
989	AgCl–TlCl	55	210.0	2651		
990	$BeCl_2$–KCl–NaCl	45–28–27	210.0	2978		
991	$FeCl_3$–K_2UCl_6	75	210.0	2985		
992	$FeCl_3$–KCl–UCl_4	NA	210.0	2985		
993	InI_2–SnI_2	74.5	210.0	2994		
994	CsN_3–$Zn(N_3)_2$	NA	210.0	3072		
995	$KClO_4$–$LiClO_4$	24	210.0	2787		
996	AgI–InI_2	20	210.0	2919		
997	Tl_2CO_3–$TlNO_3$	40	210.0	3217		
998	$AlCl_3$–KCl–$NbCl_5$	29–50–21	211.0	840		
999	AgCl–AgI	42	211.0	3170		
1000	LiOH–$NaNO_3$–NaOH	26.3–4.5–69.2	211.0	3211		
1001	Tl_2CO_3–$TlNO_3$	62	211.0	3217		
1002	$KC_2H_3O_2$–KNO_3	39	212.0	1215		
1003	$PbCl_2$–$TeCl_4$	2 APP	212.0	2709		
1004	KNO_3–$NaC_2H_3O_2$	60	212.0	2901		
1005	$NbCl_5$–$TaCl_5$	57	213.0	241		
1006	TlCl–$ZnCl_2$	48	213.0	450		
1007	KNO_2–$Sr(NO_2)_2$	73.4	213.0	2285		
1008	$Ca(NO_3)_2$–$LiNO_2$	33.3	213.0	1055		
1009	$LiC_2H_3O_2$–$NaC_2H_3O_2$	47.5	213.0	1218		
1010	Na–S	25	213.0	2631		
1011	$SeCl_4$–WCl_6	40	213.0	2718		
1012	K_2FeCl_6–K_2UCl_6	70	213.0	2985		
1013	$KAlCl_4$–K_2NbOCl_5	86.8	214.0	1048		
1014	NH_4Cl–$TaCl_5$	4 APP	214.0	1041		
1015	KNO_3–KOH	68.5	214.0	1033		
1016	$CdCl_2$–$TeCl_4$	4 APP	214.0	2709		
1017	$NaC_3H_7O_2$–NaCNS–$NaNO_2$	57.5–40–2.5	214.0	2723		
1018	$NaC_3H_7O_2$–NaCNS–$NaNO_2$	61–37–2	214.0	2723		
1019	$Ca(NO_3)_2$–$NaNO_3$	31	214.0	3003		
1020	$FeCl_3$–$ZnCl_2$	30	214.0	3138		
1021	$Ba(NO_3)_2$–KNO_3–$NaNO_3$	2.6–51.3–46.1	214.1	1872		
1022	LiCl–$SnCl_2$	15	215.0	666		
1023	KCl–$TaCl_4$	49	215.0	1027		
1024	CsCl–CuCl	23	215.0	719		
1025	$BiCl_3$–$ZnCl_2$	89.1	215.0	964	965	
1026	CuCl–$TaCl_5$	28.3 LT	215.0	889		
1027	AlI_3–NaI	48	215.0	2523		
1028	KNO_2–$Sr(NO_2)_2$	72.6	215.0	1020		
1029	$BeCl_2$–NaCl	NA	215.0	3128		
1030	AgCl–AgI	46	215.0	3171		
1031	KCl–$MgCl_2$–$ZrCl_4$	36.5–3.5–60.	216.0	1125		
1032	$BeCl_2$–NaCl	50	217.0	1941		
1033	$BiCl_3$–WCl_6	92	217.0	2054		

Locator number	System	Mol %	T, °C	References
1034	KNO_2–$Sr(NO_2)_2$	49	217.0	1020
1035	$CsNO_3$–KNO_3	39.6	217.0	1212
1036	$KC_2H_3O_2$–$RbNO_3$	68	217.0	1225
1037	$LiC_2H_3O_2$–$NaNO_3$	80	217.0	2901
1038	KNO_3–$Pb(NO_3)_2$	76.6	217.1	1943
1039	AgCl–TlCl	56.1 APP	218.0	1021
1040	$InCl_3$–$SnCl_2$	8.7	218.0	750
1041	$NaNO_2$–$Sr(NO_2)_2$	71	218.0	1020
1042	NaCNS–$NaNO_2$	41.5	218.0	2723
1043	NaCNS–$NaNO_2$	37	218.0	2723
1044	CdI_2–InI_2	3.5	218.0	2994
1045	KCl–NaCl–$ZrCl_4$	NA	218.0	3047
1046	$BiCl_3$–$PbCl_2$	88.7	219.0	1918
1047	$LiClO_4$–$NaClO_4$	78.2	219.0	913
1048	KNO_2–$Sr(NO_2)_2$	49.6	219.0	2285
1049	$CsC_2H_3O_2$–$KC_2H_3O_2$–$LiC_2H_3O_2$	13–5–82	219.0	1464
1050	$LiC_2H_3O_2$–$NaC_2H_3O_2$	69	219.0	1218
1051	$LiClO_4$–$NaClO_4$	78	219.0	2614
1052	AgCl–$TeCl_4$	2.6	219.0	2707
1053	$LiClO_4$–$NaClO_4$	78	219.0	2773
1054	$BiCl_3$–$PbCl_2$	90	219.0	3138
1055	$FeCl_3$–TlCl	40 APP	219.0	3133
1056	NaCl–$TaCl_3$	55	220.0	1019
1057	$FeCl_3$–KCl	66	220.0 ±2	2537
1058	KCl–$ZrCl_4$	42.2	220.0	83 201 794
1059	$AlCl_3$–In_2Cl_3	42 APP	220.0 APP	2392
1060	KNO_3–KOH	32	220.0	1033
1061	$CsNO_2$–$CsNO_3$–$Sr(NO_3)_2$	54.8–16.1–29	220.0	1238
1062	$KC_2H_3O_2$–KNO_3	61	220.0	1215
1063	$GaCl_3$–KCl–$MgCl_2$	44–4–52	220.0	2613
1064	$NaC_3H_7O_2$–NaCNS–$NaNO_3$	17.5–40–42.5	220.0	2723
1065	$LiC_2H_3O_2$–$Zn(C_2H_3O_2)_2$	25	220.0	2957
1066	InI_2–ZnI_2	97.0	220.0	2994
1067	$Ca(NO_3)_2$–$LiNO_3$	14±2	220.0	3003
1068	$SnBr_2$–TlBr	93.7	220.5	2748
1069	K_2CO_3–KOH–LiOH	2.1–71.8–25.6	221.0	2526
1070	$Ba(NO_2)_2$–KNO_2	31.6	221.0	2123
1071	$FeCl_3$–NH_4Cl	65	221.0	3147
1072	$GaCl_3$–NH_4Cl	35 APP	222.0	1016
1073	$KC_2H_3O_2$–$LiC_2H_3O_2$	15.5	222.0	1145
1074	$K_2Cr_2O_7$–$NaNO_3$	36.1	222.0	3190
1075	$CsNO_3$–KNO_3	40 APP	222.0	2790
1076	$Ca(NO_3)_2$–$NaNO_3$	29	222.8	1998
1077	RbCl–$SnCl_2$	51.5	223.0	286
1078	$FeCl_2$–$LaCl_3$–$SnCl_2$	6–2–92	223.0	828
1079	NaCN–NaOH	31	224.0	1101
1080	KNO_2–$NaNO_2$	35	224.0	904
1081	$HgCl_2$–WCl_6	5	224.0	2718
1082	InI_2–PbI_2	97.5	224.0	2994
1083	$TeBr_4$–$TeCl_4$	0 SER SOLID SOL	224.0	2764
1084	NaI–NaOH	19.4	224.5	1101
1085	AlI_3–LiI	48 APP	225.0 APP	2284
1086	$InCl_2$–KCl	95 APP	225.0	2767
1087	NaCl–$NaNO_3$–NaOH	1.6–70.8–27.6	225.0	2836
1088	KNO_3–$Na_2Cr_2O_7$	73.3	225.0	3190
1089	$CeCl_3$–$FeCl_2$–$SnCl_2$	2–6–92	226.0	828
1090	AgBr–TlBr	62 APP	226.0	1021
1091	KOH–LiOH	68.8	226.0	2299

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References	
1092	$Ca(NO_3)_2$–$NaNO_3$	30.3	226.0	915		
1093	$LiC_2H_3O_2$–$NaC_2H_3O_2$	81.5	226.0	1146		
1094	$NaClO_4$–$NaNO_3$	38.5	226.0	2653		
1095	InI_2–MnI_2	100 APP	226.0	2994		
1096	KOH–LiOH	71	227.0	2037	2178	2526
1097	NaCNS–$NaNO_3$	44 APP	227.0 APP	1942		
1098	HgI–HgI_2	47.5	227.0	3150		
1099	NH_4Cl–$ZnCl_2$	27	228.0	964	965	
1100	KCl–KOH–LiOH	6.0–64.0–30.0	228.0	683		
1101	$Ca(ClO_4)_2$–$LiClO_4$	13.1	228.0	325		
1102	AgBr–RbBr	68	228.0	165	803	
1103	$KC_2H_3O_2$–$NaC_2H_3O_2$	55	228.0	996	1145	
1104	$CsNO_3$–$NaNO_2$	60 APP	228.0 MIN	2923		
1105	LiBr–$LiNO_3$	23	228.0	2926		
1106	AgBr–RbBr	68	228.0	3116		
1107	HgI_2–PbI_2	23	228.0	3152		
1108	KCl–KOH–LiOH	6–64–30	228.0	3188		
1109	KCl–$PbCl_2$–$ZnCl_2$	23–14–63	229.0	682		
1110	$CdBr_2$–$KC_2H_3O_2$–$NaC_2H_3O_2$	4–48–48	229.0	2728		
1111	$CdCl_2$–$SnCl_2$	10	229.0	3138		
1112	HF–KF	48.60	229.5	567		
1113	KCl–$TaCl_3$	50 APP	230.0	1019		
1114	KCl–$ZnCl_2$	51	230.0	140	200	498
1115	$BiCl_3$–$InCl_3$	100 APP	230.0	1354		
1116	$FeCl_2$–$SnCl_2$	6	230.0	828		
1117	$Ba(NO_3)_2$–LiCl–$LiNO_3$	4.6–15.4–80	230.0	833		
1118	$Ca(NO_3)_2$–LiCl–$LiNO_3$	17.9–20.5–61.6	230.0	3248		
1119	CsBr–LiBr	48.5	230.0	2055		
1120	CsBr–Cs_2SO_4–LiBr	47.5–0 APP –52.5	230.0	2055		
1121	$NaNO_2$–NaOH	77	230.0	1033		
1122	LiOH–$NaNO_3$–NaOH	13.5–43–38.5	230.0	3211		
1123	$HgBr_2$–NaBr	91.5	232.0	343		
1124	$HgBr_2$–$PbBr_2$	95	232.0	3240		
1125	KNO_2–$NaNO_2$	35	232.0	2123		
1126	$RbNO_2$–$Sr(NO_2)_2$	78.8	232.0	1020		
1127	$Ca(NO_3)_2$–$NaNO_3$	29.8	232.0	904		
1128	NH_4Cl–$ZnCl_2$	27	232.0	3147		
1129	$CdCl_2$–$SnCl_2$	10	233.0	61	62	712
1130	$MnCl_2$–$SnCl_2$	5	233.0	61	62	712
1131	$NaNO_2$–$NaNO_3$	60	233.0	915	916	
1132	LiCl–$LiClO_4$	9.5	234.0	3009		
1133	$HfCl_4$–KCl–NaCl	NA	234.0	3047		
1134	$AlCl_3$–KCl–$ZrCl_4$	46–52–2	235.0	794		
1135	$FeCl_3$–KCl–$ZrCl_4$	44–53–3	235.0	794		
1136	KCl–$ZrCl_4$	34.5	235.0	83	201	794
1137	AgCl–$BeCl_2$	60	235.0	512		
1138	$NaNO_2$–$NaNO_3$	50	235.0	915	916	
1139	$Ca(NO_3)_2$–$LiNO_3$	15	235.0	1998		
1140	$KC_2H_3O_2$–$NaC_2H_3O_2$	53.5	235.0	1738		
1141	AlI_3–CsI–NaI	45.0–53.0–2.0	235.0	2715		
1142	AgCl–$BeCl_2$	60	235.0	3128		
1143	$FeCl_3$–NH_4Cl	38	235.0	3147		
1144	KCl–KH_2PO_4–KNO_3	7–59–34	235.0	3187		
1145	Na_2S_2–Na_2S_4	50	235.0	3219		
1146	$Ca(NO_3)_2$–$LiNO_3$	15.3	235.2	557		
1147	CsCl–CuCl	45 APP	236.0	719		
1148	$KClO_3$–$NaClO_3$	11	236.0	384	3242	
1149	CsBr–KBr–LiBr	25–19–56	236.0	2111		

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References
1150	$Ca(NO_3)_2$–$CsNO_2$	39	236.0	1209	
1151	$Ca(NO_3)_2$–$LiNO_3$	15.3	236.0	1239	
1152	$LiC_2H_3O_2$–$RbC_2H_3O_2$	88.5	236.0	952	
1153	$CsC_2H_3O_2$–$CsNO_3$–$LiC_2H_3O_2$	11.5–5–83.5	236.0	2902	
1154	$HgBr_2$–$HgSO_4$	99 GT	236.0	3177	
1155	$HgCl_2$–$HgSO_4$	99 GT	236.0	3177	
1156	$LaCl_3$–$SnCl_2$	2.3	237.0	828	
1157	$K_2Cr_2O_7$–LiCl	30.7	237.0	1938	
1158	NaOH–RbOH	75.2	237.0	512	
1159	$RbNO_2$–$Sr(NO_2)_2$	40	237.0	1020	
1160	Na_2S_4–Na_2S_5	33	237.0	3219	
1161	$CaCl_2$–$LiNO_3$	8.1	238.0	3248	
1162	AgI–KI	29.5±1	238.0	1839	
1163	$K_2CH_3O_2$–$LiC_2H_3O_2$	13	238.0	1218	
1164	As_2Se_3–Tl_2Se	21	238.0 ±3	3224	
1165	$KC_2H_3O_2$–$Mg(C_2H_3O_2)_2$	63.5	238.0	2794	
1166	LiOH–RbOH	26	238.0	2825	
1167	$AlCl_3$–$HfCl_4$–KCl	48–1–51	239.0	1124	
1168	$NaNO_2$–NaOH	27	239.0	1033	
1169	$FeCl_3$–KCl–LiCl	46–50–4	239.0 ±1	2975	
1170	K_2NbOCl_5–$KTaCl_6$	40	239.0	3025	
1171	$CeCl_3$–$SnCl_2$	2.5	240.0	828	
1172	$CoCl_2$–$SnCl_2$	3	240.0	512	
1173	AgCl–AgI–KCl	43–53–4	240.0	616	
1174	$SnCl_2$–SnS	97.2	240.0	883	
1175	AlI_3–CsI	46.5	240.0	2523	
1176	AlI_3–CsI	45 APP	240.0 APP	2284	
1177	PbS–SnS	17 APP	240.0	883	
1178	$CsC_2H_3O_2$–$LiC_2H_3O_2$	12	240.0	952	
1179	$KC_2H_3O_2$–$NaC_2H_3O_2$	61.5	240.0	1738	
1180	$CsI·AlI_3$–NaI	92 APP	240.0	2715	
1181	Na_2S–S	31.25	240.0 ±2	2932	
1182	$NaC_2H_3O_2$–$Zn(C_2H_3O_2)_2$	50	240.0	2957	
1183	CsOH–LiOH	74.5	240.0 ±5	3095	
1184	$LiC_2H_3O_2$–$LiCHO_2$	62.5	240.0	3100	
1185	$InBr_3$–$TeBr_4$	52.5	240.0	2888	
1186	$BaSO_4$–$LiNO_3$–Li_2SO_4	2.5–96–1.5	240.0	2906	
1187	KNO_3–$Na_2Cr_2O_7$	44.6	240.0	3190	
1188	AgI–KI	70	240.5	2673	
1189	NaOH–RbOH	34.1	241.0	512	
1190	NaF–Na_2TiF_6	75.1	242.0	761	
1191	$BiCl_3$–NaCl	48	242.0	2525	
1192	$HfCl_4$–KCl	62	242.0	83	
1193	$KAlCl_4$–$NbOCl_3$	97.1	242.0	2086	
1194	AgCl–$CdCl_2$–$ZnCl_2$	43–6–51	242.0	512	
1195	$PbCl_2$–$ZnCl_2$	7	242.0	35	61
1196	NaCl–$NaNO_3$–NaOH	3.6–18.3–78.1	242.0	2836	
1197	AgI–HgI_2	15	242.0	3114	
1198	HgI_2–$HgSO_4$	87 APP	242.0	3178	
1199	NH_4Cl–$TaCl_5$	70	243.0	1041	
1200	$Ca(NO_2)_2$–$CsNO_3$	39.3	243.0	1209	
1201	CuCl–$ZnCl_2$	18	243.0	3138	
1202	RbCl–WCl_5	1.5–2 APP	244.0 APP	1051	
1203	$SnCl_2$–YCl_3	89.6	244.0	1154	
1204	LiCl–$LiNO_3$	12.5	244.0	680	
1205	BiI_3–HgI_2	13.3	244.0	2452	
1206	$Ba(NO_3)_2$–$LiNO_3$	2	244.0	1161	
1207	KH_2PO_4–KNO_3	29.2	244.5	3187	

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
1208	$AgCl-NH_4Cl$	8.5	245.0	655
1209	$CsNO_3-LiC_2H_3O_2$	9.8	245.0	1164
1210	$BeCl_2-KCl-YCl_3$	38–37–25	245.0	2739
1211	$CsNO_3-LiC_2H_3O_2$	10	245.0	2902
1212	$AgI-HgI_2$	10APP	245.0	3113
1213	$NaNO_3-NaOH$	72	246.0	3211
1214	$Ca(NO_2)_2-CsNO_2$	23.1	247.0	1209 1897
1215	$FeCl_3-KCl-LiCl$	48–48–4	247.0 ±1	2966
1216	$FeCl_3-KCl-LiCl$	48–48–4	247.0 ±1	2975
1217	$KNbOCl_4-KTaCl_6$	40	247.0	3025
1218	$NaCl-NaNO_3-NaOH$	4.2–40.2–55.6	247.0	2836
1219	$NaF-SnF_2$	43	248.0	1306
1220	$NaCl-NbCl_4$	33	248.0	791
1221	$AlCl_3-TlCl$	40.6	248.0	1157
1222	$CsNO_2-Sr(NO_2)_2$	82.3 APP	248.0	1020
1223	Na_2S-S	21.0	248.0 ±2	2932
1224	$AgI-TeI_4$	16	248.0	3022
1225	$Ba(NO_3)_2-BaSO_4-LiNO_3$	1–1–98	248.0	2906
1226	$RbCl-ZnCl_2$	47.5	249.0	1918
1227	$AlCl_3-KCl-LiCl$	48–48–4	249.0 APP	2975
1228	$As_2Se_3-Tl_2Se$	72	249.0 ±3	3224
1229	$AgCl-HgCl$	57.4	250.0	655
1230	$InCl_2-TlCl$	84 APP	250.0 APP	1462
1231	$AgCl-LiCl-LiNO_3$	0.5–12–87.5	250.0	399
1232	$K_2UCl_6-LiCl-Li_2UCl_6$	34.1–26.9–39	250.0 ±2	2865
1233	$KI-KOH$	27	250.0	3193
1234	$LiNO_3-Pb(NO_3)_2$	96.5	250.2	1943
1235	$Ba(NO_3)_2-LiNO_3$	2.8	250.4	1942
1236	$KNbCl_6-NbOCl_3$	78.4	251.0	2086
1237	$AgCl-RbCl$	60	251.0	61 62 2046
1238	$SnCl_2-TlCl$	44	251.0	2061
1239	$Ba(NO_3)_2-LiNO_3$	2.8	251.0	993
1240	$LiNO_3-Sr(NO_3)_2$	98.5	251.3	1942
1241	$AgCl-CuCl$	45.7	252.0	850
1242	$LiNO_3-Li_2SO_4$	1.26	252.0	511
1243	$Ba(NO_2)_2-Ba(NO_3)_2$	94	252.0	1163
1244	$Ba(NO_3)_2-LiNO_3$	2.6	252.0	833
1245	$Mg(CHO_2)_2-NaCHO_2$	21	252.0	2794
1246	$NaF-SnF_2$	92	253.0	1306
1247	$CdBr_2-CsBr-KBr$	37–12.3–50.7	253.0	2527
1248	$CdI_2-PbBr_2-PbI_2$	12.7–56.4–30.9	253.0	1676
1249	Na_2S-S	18.6	253.0 ±2	2932
1250	$AgBr-Ag_2SO_4$	69 APP	254.0	3169
1251	$SnCl_2-TlCl$	NA	254.3	2748
1252	$CsCl-NbOCl_3$	40	255.0	1050
1253	$CsCl-TaCl_4$	52	255.0	1007
1254	$KI-MgI_2$	61	255.0	1918
1255	$CuCl-MgCl_2-PbCl_2$	57–5–38	255.0 ±3	2694
1256	NpF_4-TlF	14	255.0	2994
1257	$Ba(NO_2)_2-Ba(NO_3)_2$	94	255.0	3039
1258	$Mg(C_2H_3O_2)_2-NaC_2H_3O_2$	42.5	255.0	2794
1259	$Ba(NO_2)_2-Ca(NO_2)_2$	89	255.0	2804
1260	$Mg(C_2H_3O_2)_2-NaC_2H_3O_2$	60	256.0	2794
1261	$NaF-SnF_2$	57	258.0	1306
1262	$AgCl-CsCl$	72	258.0	719
1263	$AgCl-CuCl$	48.9	258.0	850
1264	$CuCl-PbCl_2$	NA	258.0	3150
1265	$NaNO_3-NaOH$	18.5	258.0	3211

Locator number	System	Mol %	T, °C		References	
1266	LiBr–RbBr	59	259.0	430		
1267	$NaF–SnF_2$	72	260.0	1306		
1268	$FeCl_2–InCl_3–NaCl$	10–45–45	260.0	2466		
1269	AgCl–CuCl	46	260.0	61	62	715
1270	$InCl_3–TlCl$	48	260.0	873		
1271	$InCl_3–ZnCl_2$	8	260.0	1478		
1272	$CdCl_2–TlCl–TlI$	15.3–54.2–30.5	260.0	790		
1273	KBr–LiBr–RbBr	WORK IN PROGRESS	260.0 APP	2442		
1274	$NaBr–PbBr_2–PbI_2$	11.3–53.8–34.9	260.0	1995		
1275	$PbBr_2–PbI_2$	53	260.0	948		
1276	$PbBr_2–PbI_2$	63	260.0	1995		
1277	NaBr–NaOH	77.7	260.0	512		
1278	KI–LiI	36.9	260.0	1968		
1279	$NaBO_2–NaOH$	15	260.0	1295		
1280	$ReCl_3–ReCl_5$	15	260.0	2683		
1281	$CuI–TeI_4$	23 APP	260.0	3022		
1282	$CaBr_2–LiBr–RbBr$	3.1–55.7–41.2	260.0	2770		
1283	$K_2UCl_6–Li_2UCl_6–UCl_4$	32.9–41.1–26.0	260.0 ±2	2865		
1284	$PbCl_2–ZnCl_2$	9.1	261.0	964		
1285	$CdCl_2–ZnCl_2$	0 APP	261.5	1918		
1286	$CdCl_2–ZnCl_2$	1	262.0	964	965	
1287	LiCl–LiOH	37	262.0	683		
1288	CsBr–LiBr	37.5	262.0	1121		
1289	$AlBr_3–RbBr$	49	262.0	2265		
1290	$InBr_3–RbBr$	55	262.0	3101		
1291	$FeCl_3–TlCl$	26	262.0	3133		
1292	LiCl–LiOH	37	262.0	3188		
1293	$Na_2Cr_2O_7–NaNO_3$	45.4	262.0	3190		
1294	$CsCl–ZnCl_2$	42.5	263.0	1918		
1295	$CuCl–FeCl_3$	39.6	263.0	63		
1296	$Na_2CO_3–Na_2O–NaOH$	6.5–7.4–86.1	263.0	1138		
1297	$Ba(NO_2)_2–TlNO_2$	97	263.0	910		
1298	$CuCl–FeCl_3$	NA	263.0	3138		
1299	KI–NaCNS	17.9	263.6	1940		
1300	$HgCl_2–NaCl$	86	264.0	513		
1301	AgCl–AgI	53.5	264.0	616		
1302	KI–LiCl–LiI	WORK IN PROGRESS	264.0 APP	2442		
1303	AgBr–KI	80.3	264.0	1918		
1304	AgCl–AgI	53.5	264.0	3113		
1305	$HgCl_2–NaCl$	86	264.0	3151		
1306	$K_2Cr_2O_7–KNO_3$	35.7	264.0	3190		
1307	KCl–LiCl–RbCl	WORK IN PROGRESS	265.0 APP	2442		
1308	$HgCl_2–InCl_3$	98	265.0	1478		
1309	$Ca(NO_3)_2–LiCl$	40.85	265.0	1239		
1310	KI–LiBr–LiI	WORK IN PROGRESS	265.0 APP	2442		
1311	$MnCl_2–ZnCl_2$	2	266.0	61		
1312	$Ba(NO_2)_2–KNO_2$	93.2	266.0	2123		
1313	$LiI–TeI_4$	3	266.0	3022		
1314	KBr–LiBr–NaBr–RbBr	NA	266.0	2870		
1315	$NaNO_3–NaOH$	41	266.0	3211		
1316	$LiCl–NH_4Cl$	50	267.0	61	224	
1317	$CdCl_2–NH_4Cl$	53	267.0	3250		
1318	$AlBr_3–RbBr$	48	267.0	2470		
1319	$Ca(NO_3)_2–CsNO_2$	63.5	267.0	1209		
1320	$KNO_3–Sr(NO_3)_2$	63.1	267.0	1237		
1321	$KNO_3–Sr(NO_3)_2$	71.9	267.0	1892		
1322	$TeBr_4–TeI_4$	20	267.0 SER SOLID SOL	2764		
1323	$TeBr_4–TeCl_4–TeI_4$	20–0–80	267.0 SER SOLID SOL	2764		

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References	
1324	$LiCl-NH_4Cl$	50	267.0		3147	
1325	$CdCl_2-NH_4Cl$	53	267.0		3147	
1326	$CsCl-ZnCl_2$	20.1	268.0		1918	
1327	$LiCl-LiOH$	42	268.0		1029	
1328	$KBr-LiBr-RbBr$	9–57–34	268.0		2111	
1329	$SnBr_2-TlBr$	51.9	268.0		2748	
1330	$NaBeF_3-NaPO_3$	60	270.0	APP	1412	1678
1331	$NaCl-TaCl_4$	55	270.0		1027	
1332	$NaCl-ZnCl_2$	42	270.0		121	140
1333	$CdCl_2-TlI$	27.0	270.0		790	
1334	$KCl-PbCl_2-PbI_2$	31.2–27.8–40.9	270.0		512	
1335	$Ca(ClO_4)_2-NaClO_4$	55	270.0		325	
1336	BaF_2-FeF_2	33	270.0		2991	
1337	$CaBr_2-CsBr-LiBr$	1.0–39.4–59.6	270.0		2759	
1338	$BeCl_2-CsCl$	59.4	270.0		2742	
1339	$SrCl_2-ZnCl_2$	NA	270.0		3166	
1340	$Bi-CdTe$	99.992	270.5		2618	
1341	$MgCl_2-ZnCl_2$	1	271.0		61	156
1342	$KNO_3-Sr(NO_3)_2$	85.6	271.7		1942	
1343	$InCl_3-NaCl$	49	272.0		855	
1344	$CsNO_3-Sr(NO_3)_2$	76.9	272.0		1238	
1345	$AgCl-HgCl_2$	16.5	272.0		3113	
1346	$CsCl-TaCl_3$	54	273.0		1019	
1347	$LiBr-RbBr$	55	273.0		430	
1348	$LiCl-LiOH$	34.5	274.0		1029	
1349	$NaNO_3-Pb(NO_3)_2$	84.5	274.2		1943	
1350	$CsCl-WCl_5$	2 APP	275.0		1051	
1351	$LiBr-LiOH$	55	275.0		511	
1352	$CsNO_3-Sr(NO_3)_2$	75	275.0		1998	
1353	$KNO_3-Sr(NO_3)_2$	85	275.0		1998	
1354	$CuCN-KCN$	41	275.0		3111	
1355	$LiBr-LiOH$	55	275.0		3202	
1356	$InCl_3-ZnCl_2$	4	276.0		397	
1357	$TaCl_5-TlCl$	36	276.0		2241	
1358	$CdCl_2-CdI_2-PbI_2$	33–13–54	276.0		282	
1359	$AgBr-PbBr_2$	NA	276.0		2430	
1360	$AgBr-PbBr_2$	54	276.0		512	
1361	$AgCl-PbCl_2-TlCl$	37–55–8	276.0		2651	
1362	$CsBr-LiBr$	36.5	276.0		2759	
1363	$KBr-RbNO_3$	.5	276.0		2900	
1364	$RbCl-ZnCl_2$	37	277.0		1918	
1365	$KCl-LiCl-PbCrO_4$	36.3–50–13.6	277.0		1054	
1366	$CdBr_2-KBr-PbBr_2$	24–52–24	277.0		219	
1367	$BeCl_2-RbCl$	54.8	277.0		3102	
1368	$KF-SnF_2$	45	278.0		1306	
1369	$HgCl_2-LiCl$	98.4	278.0		513	
1370	$CsI-PbCl_2-PbI_2$	23.8–28.6–47.6	278.0		2198	
1371	$KCl-KI-PbI_2$	4.8–33.2–61.9	278.0		371	
1372	$NaCl-NaNO_3-Na_2SO_4$	8.4–86.3–5.3	278.0		512	
1373	$KBr-LiBr-PbBr_2$	43–40.7–16.3	278.0		885	
1374	$Na_2MoO_4-NaNO_2$	2 APP	278.0		2729	
1375	$HgCl_2-LiCl$	98.4	278.0		3151	
1376	$NaNO_2-Na_2WO_4$	98 APP	279.0		2729	
1377	$NaCNS-RbI$	88.5	279.2		1940	
1378	$KF-SnF_2$	56	280.0		1306	
1379	$CuCl-PbCl_2$	59.1	280.0		512	
1380	$Ba(NO_3)_2-NaCl-NaNO_3$	5.8–7.9–86.2	280.0		512	
1381	$KCl-LiCl-LiOH$	11.5–45.0–43.5	280.0		683	

Locator number	System	Mol %	T, °C	References
1382	$CdBr_2$–NaBr–$PbBr_2$	13–25–62	280.0	211
1383	$CdBr_2$–KBr–$PbBr_2$	17–11–72	280.0	219
1384	NaOH–Na_2S	95.7	280.0	1420
1385	As_2Se_3–As_2Te_3	47	280.0	2642
1386	CuCN–KCN	76	280.0	3111
1387	CuCl–$PbCl_2$	51	280.0	3138
1388	KCl–KH_2PO_4	5.5	280.0	3187
1389	KCl–LiCl–LiOH	11.5–45–43.5	280.0	3188
	[See locator number 6189.]		280.4	
1390	KI–LiCl–LiF	WORK IN PROGRESS	281.0 APP	2442
1391	$HgCl_2$–$PbCl_2$	100 APP	281.0 APP	1072
1392	$NaNO_3$–TlCl	91	281.0	1170
1393	NaCl–$NaNO_2$	1.1	281.0	61
1394	CsCl–$NbCl_4$	43 APP	282.0	387
1395	NaCl–Na_2CO_3–NaOH	7.8–6.4–85.8	282.0	213
1396	KBr–LiBr–$PbBr_2$	43.6–30.8–25.6	282.0	885
1397	$PbBr_2$–$ZnBr_2$	60	282.0	2120
1398	CuCl–$ZnCl_2$	12	283.0	190
1399	KCl–LiCl–LiOH	1.5–36.5–62.0	283.0	683
1400	Na_2CO_3–NaOH	7.2	283.0 ±1	1138
1401	$CdBr_2$–$KC_2H_3O_2$	11.5	283.0	2728
1402	KCl–LiCl–LiOH	1.5–36.5–62	283.0	3188
1403	RbCl–$TaCl_4$	53 APP	284.0 APP	1007
1404	CuCl–CuI	57	284.0	61
1405	$CdBr_2$–NaBr–TlBr	25–2.5–72.5	284.0	2472
1406	$CsNO_2$–$Sr(NO_2)_2$	40 APP	284.0	1020
1407	KBr–NaCNS	9.1	284.6	1940
1408	NaCNS–RbBr	91.3	284.9	1940
1409	$CdCl_2$–CsCl–TlBr	26.6–7.6–65.8	285.0	2562
1410	CdI_2–NaCl–NaI	36.2–36.2–27.5	285.0	321
1411	LiCl–Li_2SO_4–$ZnCl_2$	23–0.44–76.5	285.0	425
1412	AgBr–KBr	68	285.0	61 62 87 112 162 376 738 803
1413	KNO_3–$NaNO_2$	45.5	285.0	2122
1414	$CsNO_3$–$RbNO_3$	14	285.0	3251
1415	ThF_4–TlF	10	285.0	2713
1416	NaOH–Na_2S	99.5	285.0	2967
1417	CsCl–$ZrCl_4$	32.8	286.0	83
1418	AgCl–$CdCl_2$–$PbCl_2$	39.7–23.7–36.6	286.0	512 665
1419	AgBr–KBr	69	286.0	61 62 87 112 162 376 738 803
1420	$PbCl_2$–$ZnCl_2$	23	286.0	3029
1421	AgI–NaI–$NaNO_3$	0.5–14–85.5	286.0	3115
1422	Ag_2SO_4–KNO_3	9.8	286.1	2646
1423	CsBr–NaCNS	7.9	286.3	1940
1424	LiBr–NaCNS	8.7	286.8	1940
1425	$Ba(NO_3)_2$–KNO_3	12.4	287.0	1942
1426	$Ba(NO_3)_2$–KNO_3	12.5	287.0	1237
1427	$Ba(NO_3)_2$–KNO_3	12.7	287.0	993
1428	$Ba(NO_3)_2$–KNO_3	13.3	287.0	1252
1429	CdI_2–NaI	53	287.0	3140
1430	NaCNS–NaI	83.7	287.5	1940
1431	$CdBr_2$–TlCl	23.8	288.0	2297
1432	$CsNO_3$–$RbNO_3$	20 APP	288.0	2790
1433	KBr–LiBr–$PbBr_2$	15–18.3–66.6	289.0	885
1434	CsBr–NaBr–$PbBr_2$	7.9–15.8–76.2	289.0	1793
1435	$CdCl_2$–NaCl–TlCl	22–4–74	290.0	480
1436	KCl–$NbCl_5$–$ZrCl_4$	51–47–2	290.0	874

Locator number	System	Mol %	*T*, °C	References
1437	$CdCl_2$–CsCl–TlCl	21.2–3.6–75.1	290.0	2562
1438	$CdCl_2$–TlCl	22.7	290.0	790
1439	$GaCl_3$–TlCl	32	290.0	1016
1440	AgCl–Ag_2SO_4–Li_2SO_4	NA	290.0	440
1441	$CdCl_2$–$CdSO_4$–TlCl	16–7.4–76.5	290.0	392
1442	AgBr–KBr	51	290.0	61 62 87 112 162 376 738 803
1443	$Ca(NO_2)_2$–$RbNO_2$	56.2	290.0	1897
1444	$CsNO_3$–$RbNO_3$	15 APP	290.0 APP	1486
1445	KCl–NaCl–$ZnSO_4$	25–20–55	290.0	2726
1446	NaI–SnI_2	22	290.0	2733
1447	LiCl–Li_2SO_4–RbCl	58.2–1.3–40.5	290.0	2763
1448	CsCl–LiCl–RbCl	5–56.5–38.5	290.0	2819
1449	KCl–NaCl–$ZnSO_4$	33–33.5–33.5	290.0	2897
1450	AgCN–KCN	14	290.0	3111
1451	AgBr–KBr	51	290.0	3116
1452	LiCl–LiOH	NA	290.0	3202
1453	NaCl–Na_2CO_3–NaOH	6–6.6–87.3	291.0	213
1454	KNO_3–$RbNO_3$	30	291.0	1082
1455	$BeCl_2$–$PbCl_2$	53	292.0	512
1456	$InCl_3$–TlCl	50	292.0	992
1457	$CdBr_2$–CsCl–TlBr	27.4–10.2–62.4	292.0	2562
1458	$CdBr_2$–KBr–NaBr	35–52–13	292.0	210
1459	Cs_2CO_3–CsOH	10.3	292.0	1024
1460	$BeCl_2$–CsCl	75.0	292.0	2742
1461	$BeCl_2$–$PbCl_2$	47	292.0	3128
1462	Li_2SO_4–$ZnCl_2$–$ZnSO_4$	0.7–90.3–9.0	293.0	425
1463	NaCl–Na_2SO_4–NaCNS	4.2–1.0–94.8	293.0	246
1464	NaOH–Na_2SO_4	95.3	293.0	1420
1465	$CdBr_2$–$Na_2C_2H_3O_2$	25	293.0	2728
1466	NaBr–$NaNO_3$	9.5	293.0	3205
1467	NaCNS–RbCl	94.4	293.4	1940
1468	$NaNO_3$–$Sr(NO_3)_2$	93.6	293.4	1942
1469	CsCl–NaCNS	5.3	293.6	1940
1470	LiCl–$ZnCl_2$	23	294.0	425
1471	LiCl–NaCNS	6	294.0	1940
1472	$Ba(NO_3)_2$–$NaNO_3$	6.4	294.0	993
1473	KCl–NaCNS	5.2	294.1	1940
1474	$NaNO_3$–$Sr(NO_3)_2$	93	294.9	1998
1475	KF–NaF–KNO_3	7–2.5–90.5	295.0	381
1476	$CdCl_2$–$PbCl_2$–PbI_2	31.5–37–31.5	295.0	282
1477	KCl–$ZnSO_4$	66.6	295.0	327
1478	$CdBr_2$–KBr–TlBr	28.2–24.3–47.4	295.0	2329
1479	$CdBr_2$–$PbBr_2$–TlBr	16–29.9–54	295.0	2469
1480	NaI–$PbBr_2$	22.2	295.0	1995
1481	$K_2S_2O_7$–V_2O_5	92	295.0	2585
1482	$Ba(NO_3)_2$–$NaNO_3$	5.8	295.0	1942
1483	$CdCl_2$–CsBr–TlBr	26.6–2.6–70.8	295.0	3004
1484	As_2S_3–In_2S_3	98 APP	295.0 ±5	2761
1485	$BeCl_2$–KCl–YCl_3	62–24–14	295.0	2739
1486	$CdCl_2$–TlBr	25.8	296.0	2297
1487	NaI–$NaNO_3$	14	296.0	1100
1488	Na_2MoO_4–$NaNO_3$	2 APP	296.0	2729
1489	NaBr–NaCNS	9.4	296.1	1940
1490	CsCl–LiCl–LiF	38–59.5–2.5	297.0	1223
1491	NaCl–$NaNO_3$	6.5	297.0	1100
1492	$CdBr_2$–KBr	36	297.0	72 210 219 923
1493	$CdBr_2$–KBr–$PbBr_2$	38–31–31	297.0	219

Table 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References			
1494	KCl–LiCl–NaCl–RbCl	18.3–50.4–8.0–23.3	297.0	3088			
1495	$FeCl_2$–$FeCl_3$	13.5	297.5	721			
1496	$CdCl_2$–CdF_2–NaF	46.3–9.3–44.4	298.0	2468			
1497	$CdCl_2$–NaF	55 APP	298.0 ±5	26			
1498	KF–KNO_3	9	298.0	381			
1499	CsCl–LiCl–$SrCl_2$	39.8–58.2–2.0	298.0	2008			
1500	KCl–$NbCl_4$	40 APP	298.0	791			
1501	$CdCl_2$–TlCl	29.5	298.0	2506			
1502	NaCl–$NaNO_3$	5	298.0	61	62		
1503	KCl–$KClO_3$–KNO_3	6.9–18.6–74.5	298.0	661			
1504	$CdBr_2$–$PbBr_2$–TlBr	17.6–72.8–9.5	298.0	2469			
1505	KBr–PbI_2	47.3	298.0	948			
1506	$Ba(NO_3)_2$–$NaNO_3$	6.4	298.0	894			
1507	$NaNO_3$–Na_2WO_4	98 APP	298.0	2729			
1508	SnI_2–SnS	85 APP	298.0	2997			
1509	$CdBr_2$–KBr	37	299.0	2071			
1510	$KBeF_3$–KPO_3	80 APP	300.0	1413			
1511	$KBeF_3$–KPO_3	80	300.0 APP	3245			
1512	$BeCl_2$–LiCl	56	300.0	512			
1513	KCl–UCl_4	57	300.0	1394			
1514	RbCl–$TaCl_3$	55	300.0	1019			
1515	$CdCl_2$–TlCl	27	300.0	711			
1516	$CdCl_2$–$ZnCl_2$	8	300.0	511			
1517	$FeCl_3$–$ZrCl_4$	85.18	300.0	794			
1518	CdI_2–TlCl	33.3	300.0	790			
1519	$ZnCl_2$–$ZnSO_4$	90	300.0	425			
1520	$CdBr_2$–KBr	37	300.0	72	210	219	923
1521	KBr–KOH	25	300.0	511			
1522	$NaNO_3$–Na_2SO_4	95.5	300.0	511			
1523	Rb_2SO_4–$RbNO_3$	1.5	300.0	2529			
1524	$CdCl_2$–In_2Cl_3	9	300.0	2705			
1525	$BeCl_2$–KCl	48	300.0	2978			
1526	$FeCl_3$–$ZrCl_4$	94	300.0 ±2	3020			
1527	$FeCl_3$–$HfCl_4$	62	300.0 ±2	3020			
1528	$BeCl_2$–KCl	47.9	300.0	3048			
1529	RbCl–TlCl–TlI	NA	300.0	2757			
1530	LiBr–RbBr	54	300.0	2770			
1531	AgBr–$TeBr_4$	20	300.0	2875			
1532	RbBr–$RbNO_3$	1	300.0	2900			
1533	$BeCl_2$–LiCl	56	300.0	9559			
1534	$BeCl_2$–TlCl	NA	300.0 APP	3128			
1535	$K_2Cr_2O_7$–$Na_2Cr_2O_7$	42.5	300.0	3190			
1536	KBr–KOH	25	300.0	3193			
1537	Na_2CrO_4–$NaNO_3$	5 APP	300.0	3210			
1538	NaCl–NaCNS	6	300.3	1940			
1539	$CdCl_2$–CdF_2–NaF	49.4–3.8–46.7	301.0	2468			
1540	Na_2CrO_4–$NaNO_3$	5 APP	301.0	3209			
1541	CsCl–$HfCl_4$	34.9	302.0	83			
1542	NaCl–NaCNS	4.8	302.0	246	291		
1543	$CdBr_2$–$PbBr_2$–TlBr	27.4–43.6–29	302.0	2469			
1544	$BaSO_4$–LiCl–RbCl	1–57–42	302.0	2793			
1545	$CdBr_2$–$PbBr_2$–TlBr	20.3–19.6–60.1	303.0	2469			
1546	KBr–K_2CrO_4–LiBr	15.3–8.9–75.8	303.0	1938			
1547	$Na_2Cr_2O_7$–$Rb_2Cr_2O_7$	78	303.0	1112			
1548	$NaBF_4$–NaF	61.1	304.0	1039			
1549	NaF–$NaNO_3$	3.5	304.0	381			
1550	AgCl–Ag_2SO_4	69.3	304.0	208	440		
1551	$CdSO_4$–TlCl–Tl_2SO_4	19.7–72.6–7.6	304.0	392			

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
1552	$KBr-NaBr-PbBr_2$	7.9–16.7–75.4	304.0	2498
1553	$CdBr_2-KBr$	38	304.0	965
1554	$AgCl-Ag_2SO_4$	69	304.0	3124
1555	$CuCl-FeCl_3$	NA	304.0	3138
1556	$KF-SnF_2$	77	305.0	1306
1557	$CuCl_2-FeCl_2$	48.5	305.0	1830
1558	$CdBr_2-KBr$	37	305.0	1918
1559	CdI_2-SnI_2	30	305.0	2692
1560	$As_2S_3-Na_2S$	85 APP	305.0 ±5	3091
1561	$AgCl-Ag_2SO_4-CdSO_4$	68–1–31 APP	305.0	3112
1562	KNO_3-TlCl	88.9	305.5	1170
1563	$BeCl_2-BeF_2$	27.5	306.0	1918
1564	$CsCl-LiCl$	40.5	306.0	1223
1565	$AgCl-KCl$	70	306.0	61 62 85 87 110 376 477 616 675 738
1566	$PbCl_2-PbI_2$	24	306.0	965
1567	$KBr-PbI_2$	27.6	306.0	948
1568	$KOH-RbOH$	65.6	306.0	512
1569	$LiF-NaCNS$	0.6	306.5	1940
1570	$KBF_3OH-KBF_4$	91.5	307.0	1062
1571	$CsCl-PbI_2$	23.2	307.0	2198
1572	$CsBr-KBr-PbBr_2$	11–4.6–84.3	307.0	1796
1573	$NaI-PbI_2-TlI$	9.1–65.3–25.6	307.0	1126
1574	$BaCl_2-LiCl-RbCl$	1.2–91.6–7.2	307.0	3074
1575	$KBr-LiCl-PbBr_2$	13.7–14.6–71.7	308.0	995
1576	$CdBr_2-CsBr-TlBr$	31.6–9.2–59.2	308.0	2561 2562
1577	$KF-SnF_2$	92	309.0	1306
1578	$NaBr-PbBr_2-TlBr$	14–80–6	309.0	52
1579	$KI-PbBr_2$	44.9	309.0	948
1580	$K_2CO_3-KOH-LiOH$	3.1–30.9–66	309.0	2526
1581	$BeCl_2-LiCl$	51 APP	310.0 APP	2507
1582	$AgCl-PbCl_2$	NA	310.0	2430
1583	$BeCl_2-TlCl$	55	310.0	512
1584	$KBr-KCl-LiBr-LiCl$	28.5–9.5–46.5–15.5	310.0	949
1585	$Li_2SO_4-ZnCl_2$	1	310.0	425
1586	$KCl-K_2CrO_4-KNO_3$	2.7–1–96.3	310.0	3247
1587	$Ba(ClO_4)_2-NaClO_4$	43	310.0	196
1588	$CdBr_2-TlBr$	25	310.0	788
1589	$BeCl_2-YCl_3$	95–5	310.0	2739
1590	$NaCl-ZrCl_4$	35.4	311.0	51 83 779
1591	$AgCl-PbCl_2$	61.5	311.5	1818
1592	$CsAlCl_4-CsCl-Cs_2NbOCl_5$	75–10–15	312.0	1048
1593	$CoCl_2-ZnCl_2$	7.3	312.0 ±1	611
1594	$FeCl_3-InCl_3$	3	312.0	1354
1595	$CsI-PbCl_2-PbI_2$	18.2–61.4–20.4	312.0	2198
1596	$CsCl-LiBO_2-LiCl$	43.5–0.5–56	312.0	2291
1597	$CaCl_2-CsCl-LiCl$	0.5–38.7–60.8	312.0	2759
1598	$CaBr_2-KBr-LiBr$	6.5–41.5–52	312.0	2818
1599	$CdCl_2-NH_4Cl$	79	312.0	3147
1600	$Ba(ClO_4)_2-NaClO_4$	40	313.0	1116
1601	$CdBr_2-TlBr$	25	313.0	2071
1602	$CsBr-LiBr$	50.0	313.0	2759
1603	$CsCl-LiCl$	42.5	314.0	363 2291
1604	$CuCl-NaCl$	73	314.0	61 62
1605	$CuCl-NaCl$	74	314.0	64
1606	$CuCl-NaCl$	75	314.0	715
1607	$NaCl-ZrCl_4$	37.6	314.0	201
1608	$AgCl-PbCl_2$	60.02	314.0	30 62 169 195 208 738

Table 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
1609	KBr–KCl–LiBr–LiCl	21.3–37.7–34.8–6.1	314.0	949
1610	KBr–LiCl–$PbBr_2$	42.6–24–33.3	314.0	995
1611	NaCl–NaOH	6.3	314.0	1199
1612	$Ba(ClO_4)_2$–$Ca(ClO_4)_2$	43	314.0	1116
1613	KNO_3–K_2WO_4	94.2	314.0	2295
1614	AgCl–$PbCl_2$	40	314.0	2651
1615	AgCl–KCl	72	314.0	3090
1616	LiCl–RbCl	58.5	314.0	2763
1617	$AlCl_3$–RbCl	46 APP	315.0 APP	2284
1618	$SnCl_2$–TlCl	24	315.0	2061
1619	AgBr–KCl	76	315.0 APP	1379
1620	AgCl–KCl–KNO_3	0.39–5.61–94.0	315.0	376
1621	$CdBr_2$–CsI	45.5	315.0	1010
1622	Cs_2CrO_7–$Na_2Cr_2O_7$	21	315.0	1112
1623	TlF–YF_3	95	315.0	2958
1624	CsCl–LiCl	39	315.0	2759
1625	CuCl–NaCl	75	316.0	2254
1626	$InCl_3$–NaCl	78	316.0	397 1193
1627	$InCl_3$–KCl	47.5	316.0	992
1628	$CdBr_2$–CsCl–TlBr	21.9–32.9–45.1	316.0	2562
1629	KBr–LiCl–$PbBr_2$	45.4–31.9–22.7	316.0	995
1630	TlCl–TlI	52.5	316.0	790
1631	$CdSO_4$–TlCl	52.7	316.0	392
1632	KNO_3–K_2WO_4	92 APP	316.0	2729
1633	TlCl–TlI	52	316.0	2757
1634	$Na_2Cr_2O_7$–$Rb_2Cr_2O_7$	25	317.0	1112
1635	LiCl–RbCl	58.3	318.0	160
1636	AgCl–KCl	70	318.0	61 62 85 87 110 376 477 616 675 738
1637	CsCl–$PbBr_2$	12.2	318.0	1994
1638	Li_2CrO_4–LiOH	29	318.0	942
1639	K_2MoO_4–KNO_3	8 APP	318.0	2729
1640	KBr–LiBr	60	318.0	2926
1641	KF–NaF–TiF_4	18.3–30.2–51.5	318.0	3028
1642	KCl–KNO_3–K_2SO_4	16.75–82.25–1	318.0	3196
1643	AgCl–KCl	70	319.0	2046
1644	$Ca(NO_3)_2$–K_2CrO_4–KNO_3	0.13–0.7–99.18	319.0	546
1645	BeF_2–NaF–ThF_4	43–55–2	320.0	1260
1646	$BaCl_2$–KCl–LiCl	5.43–40.92–53.65	320.0	1166
1647	$BaCl_2$–KCl–LiCl	6.38–39.36–54.26	320.0	128
1648	KCl–LiCl–$PbCl_2$	39.2–33.6–27.1	320.0	884
1649	$CdCl_2$–KCl–$PbCl_2$	42.5–19–38.5	320.0	322 394
1650	$CdCl_2$–KCl–$PbCl_2$	43–22–34.8 APP	320.0	1147
1651	$CuCl_2$–KCl	38	320.0	38
1652	$CsAlCl_4$–Cs_2NbOCl_5	79.5	320.0	1048
1653	AgCl–KBr	75	320.0 APP	1379
1654	KCl–LiBr–NaBr	38–55–7	320.0	949
1655	$CaCl_2$–$LiNO_3$	40.3	320.0	3248
1656	KCl–KNO_3	6	320.0	341 376 661
1657	$CdBr_2$–CsBr–NaBr	53.8–20–26.1	320.0	1854
1658	$CdBr_2$–RbBr	40	320.0	2071
1659	KI–$PbBr_2$	13.1	320.0	948
1660	KBr–KNO_3	9.5	320.0	3232
1661	$Ca(NO_2)_2$–$CsNO_3$	84	320.0	1209
1662	KCl–LiCl–Li_2CrO_4	33.2–38.0–28.8	320.0	2989
1663	KBr–LiBr–Li_2CrO_4	28.2–56.5–15.3	320.0	2989
1664	KCl–$ThCl_4$–UCl_4	NA	320.0 ±2	3097
1665	CuBr–$TeBr_4$	18 APP	320.0	2875

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
1666	AgI–TlI	30	320.0	3117
1667	Ag_2SO_4–TlI	72	320.0	3117
1668	KCl–KNO_3	6	320.0	3187
1669	$CdCl_2$–LiCl–$PbCl_2$	31.4–18.2–50.5	321.0	870
1670	KCl–LiCl–$PbCl_2$	43.2–42.6–14.2	321.0	884
1671	KI–PbI_2	31.9	321.0	948
1672	KI–PbI_2–TlI	4.5–81–14.5	321.0	2857
1673	KBF_3OH–KF	98.7	322.0	1062
1674	CuCl–NaCl	73	322.0	640
1675	LiBr–$PbBr_2$	20.9	322.0	835 885 2027
1676	$CdBr_2$–RbBr	54	322.0	2071
1677	SbI_3–Sb_2S_3	24	322.0	1904
1678	AgI–ZnI_2	53	322.0	2816
1679	KBr–LiBr	38.5	322.5	985
1680	CsCl–LiCl	40.7	323.0	160
1681	$CdCl_2$–NaCl–$PbCl_2$	36–18–46	323.0	786
1682	CsBr–$PbBr_2$	19.8	323.0	1994
1683	KNO_2–KNO_3	20	323.0	917
1684	KI–PbI_2–TlI	6–62–32	323.0	2857
1685	$CdCl_2$–KCl–LiCl	22.7–46.6–30.7	324.0	880
1686	$CdCl_2$–CsCl–TlBr	20.5–26.5–53	324.0	2562
1687	KCl–LiCl–Li_2SO_4	38.9–57.1–3.9	324.0	133
1688	KBr–LiBr–NaBr	35–57.5–7.5	324.0	831
1689	NaBr–$PbBr_2$	17.7	324.0	211
1690	NaBr–$PbBr_2$	18	324.0	52 285
1691	$CdCl_2$–CsBr–TlBr	25.8–25.2–49.0	324.0	3004
1692	CuCN–KCN	58	324.0 APP	3111
1693	$CaCrO_4$–KNO_3	1 APP	324.0	3174
1694	$NaBF_4$–NaF	60 APP	325.0 APP	217
1695	AgCl–KCl	75	325.0	61 62 85 87 110 376 477 616 675 738
1696	$CuCl_2$–KCl	30	325.0	38
1697	$KNbCl_6$–$KNbOCl_4$	84.2	325.0	2086
1698	B_2O_3–ThO_2	98 LT	325.0 LT	937
1699	Cs_2CrO_4–$K_2Cr_2O_7$	50	325.0	2262
1700	NaCl–$PuCl_3$–$ThCl_4$	46.5–18.5–35	325.0	3105
1701	LiCl–$PbCl_2$–$ThCl_4$	40–36–24	325.0 ±2	3043
1702	$PbBr_2$–RbBr	88 APP	325.0	3163
1703	K_2CrO_4–KNO_3	1 APP	325.0	3174
1704	KCl–$PbCl_2$–$ZnCl_2$	52–18–30	326.0	682
1705	$CdCl_2$–TlCl	22.7	326.0	392 480
1706	$PbCl_2$–PbI_2	35	326.0	167 511
1707	CsCl–LiCl–Li_2SO_4	44.3–53.9–1.7	326.0	363
1708	LiCl–Li_2SO_4–TlCl	34.2–0.5–65.3	326.0	356
1709	$CdBr_2$–PbI_2	37	326.0	1676
1710	K_2CO_3–KNO_3	3.7	326.0	3186
1711	BeF_2–KF	58	327.0	1179
1712	$BeCl_2$–$CdCl_2$	85	327.0	512
1713	KCl–LiBr	39	327.0	836 949
1714	NaI–PbI_2–TlI	5.5–43.9–48.9	327.0	1126
1715	$Cs_2Cr_2O_7$–$Na_2Cr_2O_7$	74	327.0	1112
1716	$BeCl_2$–$CdCl_2$	85	327.0	3128
1717	KCl–LiCl–$SrCl_2$	40–51.9–8.1	328.0	1274
1718	$CdCl_2$–KCl–$PbCl_2$	21.5–51.8–26.7	328.0	322 394
1719	$MnCl_2$–TlCl	21.5	328.0	1077
1720	AgBr–$PbCl_2$	65 APP	328.0	2430
1721	$CaSO_4$–KCl–LiCl	4.9–38.1–57	328.0	2242
1722	$PbCl_2$–$ThCl_4$–UCl_4	56–25–19	328.0	2886

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
1723	KNO_3–TlBr	97.2	328.5	1170
1724	AgCl–$InCl_3$	32	329.0	397
1725	$CdBr_2$–NaBr–TlBr	42.8–18.6–38.6	329.0	2472
1726	$CdBr_2$–CsI	64.6	329.0	1010
1727	KCl–LiCl–$NbCl_3$	42.3–55.5–2.2	330.0	1479
1728	$HfCl_4$–NaCl	59.4	330.0	83
1729	CuCl–HgCl	56.2	330.0	655
1730	$PbCl_2$–UCl_4	65	330.0	2214
1731	NaCl–$PbCl_2$–PbI_2	9.5–25.7–64.7	330.0	323
1732	Li_2SO_4–TlCl–Tl_2SO_4	1.8–79.5–18.7	330.0	356
1733	KBr–$PbBr_2$	16	330.0	782
1734	KBr–$PbBr_2$	17.3	330.0	948
1735	$AlBr_3$–CsBr	47.5	330.0	2470
1736	$CdBr_2$–CsBr–TlBr	26.6–31.6–41.8	330.0	2561 2562
1737	Na_2CO_3–Na_2SO_4	66.5 APP	330.0 APP	857
1738	NaCl–$PuCl_3$–$ThCl_4$	58.5–18.5–23	330.0	3105
1739	NaCl–$ThCl_4$–UCl_3	46–36.5–17.5	330.0 ±2	2805
1740	NaCl–$PbCl_2$–$ThCl_4$	62–13.5–24.5	330.0	2889
1741	KNO_3–TlBr	98 APP	330.0	3195
1742	$AlCl_3$–$NaNbOCl_4$	30.2	331.0	2086
1743	$AlBr_3$–CsBr	57	331.0	2265
1744	$CaCl_2$–KCl–LiCl	5.3–44.2–50.5	332.0	2119
1745	CsCl–LiCl	41.5	332.0	363
1746	CsCl–LiCl	42	332.0	159
1747	$CdCl_2$–KCl–$PbCl_2$	20.7–52–27.3 APP	332.0	1147
1748	KCl–LiCl–LiF–NaCl	36.8–51.6–3.8–7.8	332.0	2658
1749	KCl–$ThCl_4$–UCl_4	NA	332.0 ±2	3097
1750	$PbCl_2$–$ThCl_4$–UCl_4	64–12–24	332.0	2886
1751	KBr–$MgBr_2$	65	332.5 ±1.5	512
1752	$CdCl_2$–KCl–$PbCl_2$	26.6–33.6–39.7 APP	333.0	1147
1753	LiCl–LiH–LiI	27.5–14.5–58	333.0	2442
1754	KNO_3–K_2SO_4	97.5	333.0	3197
1755	$AlCl_3$–$KNbOCl_4$	40 APP	334.0 APP	2086
1756	TlCl–$ZnCl_2$	78	334.0	450
1757	KBr–LiBr	38	334.0	831 836 2027
1758	KBr–NaBr–$PbBr_2$	37.7–11.3–50.9	334.0	2498
1759	KBr–$MgBr_2$	NA	334.0	3135
1760	KNO_3–K_2SO_4	98.8	334.0	3196
1761	KCl–UCl_4	55	334.9	3246
1762	$FeCl_2$–KCl–$NdCl_3$	40.7–56–3.3	335.0	2497
1763	$AlCl_3$–CsCl	43 APP	335.0 APP	2284
1764	KBr–$PbBr_2$	10	335.0	781
1765	LiCl–NaCl–UCl_4	63.5–14.5–22	335.0 ±2	3106
1766	KCl–LiCl–UCl_3	58.5–29.5–12	335.0 ±2	3106
1767	KCl–$ThCl_4$–UCl_3	60.5–29.5–10	335.0	2805
1768	KCl–K_2UCl_6–LiCl	55.6–14.7–29.7	335.0 ±2	2865
1769	CsI–TlCl	21	336.0	2240
1770	AgCl–Ag_2CrO_4–Li_2CrO_4	74.4–20.5–5.0	336.0	764
1771	$CaCrO_4$–KCl–LiCl	2–40.8–57.2	336.0	1696
1772	CsBr–KBr–$PbBr_2$	7.6–39.7–52.7	336.0	1796
1773	NaCl–$ThCl_4$–UCl_3	56.5–23.5–20	336.0 ±2	2805
1774	CsOH–CsF	92.6	337.0	1959
1775	$CdCl_2$–KCl–LiCl	2.6–42–55.4	337.0	880
1776	CsCl–$SnCl_2$	61.5	337.0	286
1777	KBr–KNO_3	1	337.0	2900
1778	KBr–KNO_3–TlBr	4.5–93.5–1	337.0	3195
1779	CaF_2–KCl–LiCl	1.8–42.2–56	338.0	852
1780	$PbCl_2$–PbI_2	64	338.0	167 511

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References
1781	Li_2SO_4–TlCl–Tl_2SO_4	1.2–80.9–17.8	338.0	356	
1782	$CdBr_2$–KBr	54	338.0	2071	
1783	$PbCl_2$–PbI_2	NA	338.0	3150	
1784	BeF_2–NaF–UF_4	43.5–56–.5	339.0	856	
1785	K_2CrO_4–Li_2CrO_4–LiOH	10–43–47	339.0	942	
1786	BeF_2–NaF	44.3	340.0	2394	
1787	BeF_2–NaF	64	340.0	1042	
1788	BeF_2–KF	62	340.0 ±5	310	
1789	$CaCl_2$–KCl–LiCl	5.8–43.3–50.9	340.0	98	
1790	$CeCl_3$–NaCl–$ThCl_4$	2.6–67.4–30.0	340.0	54	
1791	$CdCl_2$–KCl–$PbCl_2$	25.2–53.1–21.3 APP	340.0	1147	
1792	$FeCl_2$–KCl	39.8	340.0	2497	
1793	AlOCl–$NbOCl_3$	37.4	340.0	2565	
1794	TlCl–$ZnCl_2$	73	340.0	450	
1795	$CdCl_2$–$PbBr_2$	20	340.0	284	
1796	KBr–LiCl–NaCl	35–57.5–7.5	340.0	949	
1797	CsCl–TlI	33	340.0	2240	
1798	$CdBr_2$–$PbBr_2$	15	340.0	219	284 1676
1799	NaI–PbI_2–TlI	7.2–20.5–72.2	340.0	1126	
1800	LiCl–NaCl–UCl_4	36–22.5–41.5	340.0 ±2	3106	
1801	NaSb–Na_3Sb–Na_3SbS_3	NA	340.0	3225	
1802	KCl–$ThCl_4$–UCl_4	NA	340.0 ±2	3097	
1803	$BeCl_2$–KCl–YCl_3	11–46–43	340.0	2739	
1804	$CoBr_2$–$TeBr_4$	15	340.0	2841	
1805	$MgCl_2$–$PbCl_2$–UCl_4	4–61–35	340.0 ±2	2868	
1806	LiCl–LiF–LiI	29.1–11.7–59.2	340.9	2711	
1807	CsI–PbI_2	19.8	341.0	2198	
1808	LiCl–LiF–LiI	29.1–11.7–59.2	341.1	2442	
1809	KCl–UCl_4	50	341.1	4446	
1810	$BaCl_2$–$CdCl_2$–KCl–LiCl–NaCl	1.4–52.7–17.9–9.4–8	342.0 ±3	932	
1811	LiCl–TlCl	38	342.0	710	
1812	$CaCrO_4$–KCl–LiCl	2.8–39.8–57.3 APP	342.0	1458	
1813	LiBr–TlBr	41.5	342.0	887	
1814	NaCl–$ThCl_4$–UCl_4	NA	342.0 ±2	3097	
1815	KI–PbI_2–TlI	11–52–37	342.0	2857	
1816	NaCl–$PbCl_2$–$ThCl_4$	46–18–36	342.0	2889	
1817	BeF_2–NaF	44	343.0	1042	
1818	$AlCl_3$–CsCl–$TaCl_5$	35–57.9–7.1	343.0	240	
1819	KCl–$LiBO_2$–LiCl	41–1.5–57.5	343.0	2291	
1820	NaBr–$PbBr_2$–TlBr	8–45–47	343.0	52	
1821	$AlCl_3$–CsCl	45.7	344.0	240	
1822	$CdCl_2$–CsCl–TlBr	11.7–36.9–51.4	344.0	2562	
1823	$CdCl_2$–CdI_2–NaCl	24.4–63.4–12.2	344.0	321	
1824	$CdBr_2$–$TeBr_4$	45 APP	344.0	2841	
1825	BeF_2–NaF–UF_4	56–43.5–.5	345.0	856	
1826	$CeCl_3$–NaCl–$ThCl_4$	4.0–60.6–35.4	345.0	54	
1827	KCl–$KClO_3$	13.1	345.0	661	
1828	$CdBr_2$–KBr	54	345.0	1918	
1829	$CdBr_2$–KBr	54.5	345.0	965	
1830	KCl–$ThCl_4$–UCl_3	45–30–25	345.0	2805	
1831	CuCN–NaCN	74	345.0	3111	
1832	BeF_2–KF	72	346.0	1179	
1833	KCl–LiCl–LiF	40.5–56–3.5	346.0	907	
1834	KCl–LiCl–NaCl	36–55–9	346.0	2	133 1480
1835	NaCl–$ThCl_4$	70.3	346.0	54	
1836	$CdCl_2$–CsCl–TlCl	15.6–22–62.4	346.0	2562	
1837	KCl–K_2CrO_4–Li_2CrO_4	29.9–11.7–58.4	346.0	2989	
1838	$CdCl_2$–KCl–LiCl	11.7–44.7–43.6	347.0	880	

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
1839	$CsNO_3$–TlCl	70.6	347.0	1170
1840	KCl–LiCl–TeO_2	29.2–53.8–16.9	347.0	335
1841	KCl–LiCl	42	348.0	846
1842	NaCl–$PbCl_2$–PbI_2	7.7–50.0–42.3	348.0	323
1843	KBr–LiBr	40	348.0	61 62
1844	KBr–LiBr	NA	348.0	3135
1845	$PbBr_2$–PbF_2	92.5	349.0	802
1846	NaCl–$ThCl_4$	64.3	349.0	54
1847	KBr–$PbBr_2$	12.2	349.0	2027
1848	$BeCl_2$–RbCl	83.6	349.0	3102
1849	BeF_2–LiF–UF_4	51.5–48–0.5	350.0	58
1850	CdF_2–CdI_2	82.5	350.0 APP	1918
1851	KCl–$MnCl_2$–NaCl	28.7–45–26.3	350.0	2524
1852	$FeCl_2$–KCl	40.7	350.0 ±2	1896
1853	$CdCl_2$–CsCl–TlCl	9.3–30.6–60.1	350.0	2562
1854	CsCl–$SnCl_2$	63	350.0	834
1855	$BeCl_2$–TlCl	85	350.0	512
1856	$CdBr_2$–CsCl–TlBr	53.8–24.6–21.5	350.0	2562
1857	LiCl–Li_2CrO_4	45	350.0	764
1858	LiCl–Li_2CrO_4	83	350.0	943
1859	KBr–$PbBr_2$–TlBr	3–45–52	350.0	782
1860	$CdBr_2$–CsBr–TlBr	52.7–18.3–29	350.0	2561 2562
1861	AgBr–CuBr	58	350.0	864
1862	K_2CO_3–Li_2CO_3–LiOH	28.2–12.7–59.1	350.0	2526
1863	KCl–LiCl–NaCl	38–53.5–8.5	350.0	2843
1864	LiCl–UCl_3–UF_4	43–25–32	350.0 ±2	2830
1865	KCl–$PbCl_2$–$ThCl_4$	37.5–35–27.5	350.0	2889
1866	$BeCl_2$–TlCl	NA	350.0 APP	3128
1867	$CoCl_2$–KCl	43.5	351.0	120 503
1868	$FeCl_2$–KCl	38.2	351.0 ±1	574
1869	KCl–LiCl	42	352.0	926
1870	KCl–LiCl	43	352.0	852
1871	$CdCl_2$–$CdSO_4$–KCl	47.6–13.7–38.7	352.0	348
1872	$PbBr_2$–TlBr	89.5	352.0	52 782
1873	$Cs_2Cr_2O_7$–$Rb_2Cr_2O_7$	50	352.0	2262
1874	NaCl–$PbCl_2$–TlCl	9.2–41.3–49.5	353.0	2472
1875	KCl–LiCl	(40–43)	353.5 ±5.5	1 2 11 56 61 62 67 68 74 89 92 93 97 102 108 121 122 123 133 160 164 249 338 424 431 477 486 494 507 683 688
1876	KCl–LiCl	42	354.0	836 847 884
1877	KCl–LiCl	42.5	354.0	907
1878	KCl–LiCl–$PbCl_2$	18.6–26.3–55	354.0	884
1879	$CdCl_2$–$CdSO_4$–KCl	28.7–6.4–64.9	354.0	348
1880	AgCl–Ag_2CrO_4	59.6	354.0	943
1881	AgCl–Ag_2CrO_4	73	354.0	764
1882	KBr–$PbBr_2$	44	354.0	782
1883	KBr–$PbBr_2$	48.1	354.0	948
1884	$Ca(NO_2)_2$–$CsNO_2$	73.9	354.0	1209 1897
1885	KCl–LiCl	42	354.0	2843
1886	$CdCl_2$–KCl–NaCl	59.6–22.2–18.2	355.0 ±1	932
1887	$FeCl_2$–KCl–$NdCl_3$	29.3–66.7–4	355.0	2497
1888	KBr–$PbBr_2$	46	355.0	781
1889	$BeCl_2$–$CaCl_2$	85	355.0	3128
1890	BeF_2–LiF	52	356.0	149 429 810
1891	BeF_2–LiF–ThF_4	51.5–47.0–1.5	356.0	510

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C		References		
1892	$NaCl-ThCl_4-UCl_4$	NA	356.0 ±2	3097			
1893	$TeBr_4-Tl_2TeBr_4$	10	356.0	2875			
1894	LiBr–LiCl–LiI	NA	357.0 APP	2442			
1895	$KCl-PbCl_2-ThCl_4$	21.5–38.5–40	357.0	2889			
1896	$CdCl_2-NaCl-TlCl$	60–21–19	358.0	480			
1897	$CeCl_3-NaCl-ThCl_4$	7.0–62.0–31.0	358.0	54			
1898	$KCl-NaCl-YCl_3$	3–42–55	358.0	1208			
1899	$CdCl_2-CdSO_4-NaCl$	32–13.5–54.5	358.0	304			
1900	$TlCl-Tl_2SO_4$	77.3	358.0	392			
1901	$KI-PbI_2$	52.9	358.0	948			
1902	K_2CrO_4-KOH	8.1	358.0	942			
1903	$CdCl_2-CsBr-TlBr$	14.3–40.0–45.7	358.0	3004			
1904	$TlCl-Tl_2SO_4$	54	358.0	3124			
1905	$CdCl_2-CdSO_4-NaCl$	32–13–55	358.0	3142			
1906	$CdCl_2-CdI_2$	30	359.0	61	282	321	1676
1907	$NaBr-PbI_2$	20.6	359.0	1995			
1908	$CdCl_2-KCl-LiCl$	37.9–28.9–33.1	360.0	880			
1909	$NaCl-YCl_3$	45	360.0	853			
1910	$CuCl_2-KCl$	54	360.0	38			
1911	$MgCl_2-TlCl$	27.5	360.0	512			
1912	KBr–LiCl	39	360.0	836	949		
1913	CdI_2-PbI_2	70	360.0	1676			
1914	$KI-ZnSO_4$	65.8	360.0	3249			
1915	$CdI_2-CsI-NaI$	61.2–25–13.7	360.0	1795			
1916	K_2CrO_4-KOH	7.8	360.0	1033			
1917	$Na_2S-Na_3Sb-Na_3SbS_3$	NA	360.0	3225			
1918	LiOH–RbOH	70.5	360.0	2825			
1919	$KCl-PbCl_2-ThCl_4$	48–29–23	360.0	2889			
1920	$BiCl_3-TlCl$	12.5	360.0	3133			
1921	$CoCl_2-KCl$	NA	360.0	3143			
1922	$MgCl_2-TlCl$	27.5	361.0	1918			
1923	$CdBr_2-CsCl$	68.1	361.0	1008			
1924	$NaCl-ThCl_4$	69.5	361.0	2745			
1925	$CdCl_2-KCl-PbCl_2$	29–59–12	362.0	322	394		
1926	$CdCl_2-CsCl-PbCl_2$	46.7–13.1–40.2	362.0	1102			
1927	$CoCl_2-CuCl_2$	25.6	362.0	1830			
1928	$CsCl-CsNO_3$	30.4	362.0	357			
1929	$Ca(ClO_4)_2-KClO_4$	80	362.0	1116			
1930	$CdBr_2-CsBr-NaBr$	37.9–52.4–9.6	362.0	1854			
1931	$CdI_2-CsI-KI$	23–60.8–16.2	363.0	1794			
1932	$BeCl_2-KCl$	76	363.0	2978			
1933	$BeCl_2-KCl$	75.7	363.0	3048			
1934	$CdBr_2-ZnBr_2$	21.2	364.0	176			
1935	$BeF_2-NaF-ThF_4$	55–43–2	365.0	1260			
1936	$KCl-NbOCl_3$	30	365.0	1050			
1937	$CuCl-InCl_3$	10	365.0	992			
1938	$PbCl_2-UCl_4$	67.9	365.0	2015			
1939	$CdBr_2-CsBr-TlBr$	16.3–38.4–45.3	365.0	2561	2562		
1940	$LiCl-NaCl-UCl_3$	46–28.5–25.5	365.0	2822			
1941	CsCl–CsF–CsI	34–32–34	365.0	2832			
1942	$FeCl_2-InCl_3-NaCl$	22–20–58	366.0	2466			
1943	$NaCl-ThCl_4$	72	366.0	1049			
1944	$KCl-NbCl_4$	52	366.0	240	798		
1945	$TlCl-TlCl_3$	74.4	366.0	1057			
1946	$KCl-K_2Cr_2O_7$	27.5	366.0	675			
1947	K_2CO_3-KOH	9.3	366.0	2526			
1948	BeF_2-CsF	77.5	367.0	1986			
1949	LiBr–LiF–LiI	NA	367.0 APP	2442			

Table 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References			
1950	$CdBr_2$–$PbCl_2$	28	367.0	284			
1951	$CdBr_2$–NaBr	46	367.0	210	211	326	787
1952	$CdBr_2$–NaBr	47	367.0	787	923		
1953	$CdBr_2$–NaBr	46	367.0	2771			
1954	CsI–NaI–TlI	28–18–54	367.0	2754			
1955	RbCl–UCl_4	53	367.8	3246			
1956	$CoCl_2$–NaCl	39.5	368.0	120	199	309	503
1957	KCl–$NbCl_5$	56 APP	368.0	1280			
1958	$CdCl_2$–RbCl	72.5	368.0	512			
1959	$PbCl_2$–TlCl	42.9	368.0	512			
1960	$ZrCl_4$–ZrI_4	58 APP	368.0 APP	796			
1961	KCl–$K_2Cr_2O_7$	25	368.0	1018			
1962	KBr–$PbBr_2$	45.5	368.0	2027			
1963	$AgVO_3$–$TlVO_3$	16.5	368.0	2976			
1964	LiCl–LiI	34.2	368.3	2442			
1965	AgCl–$InCl_3$	77	369.0	397			
1966	$MnCl_2$–NaCl–NaF	61–11–28	370.0	2595			
1967	$CoCl_2$–NaCl	36	370.0	120	199	309	503
1968	CsCl–NaCl–$PbCl_2$	14.9–19.8–65.3	370.0	900			
1969	$DyCl_2$–NaCl	59.7	370.0 ±2	1814			
1970	$FeCl_2$–NaCl	44	370.0	2497			
1971	KCl–NaCl–$SmCl_3$	10–36–54	370.0	1186			
1972	KCl–NaCl–$SmCl_3$	14.3–32.5–53.2	370.0	1186			
1973	KCl–$TaCl_5$	52	370.0	240	797	798	
1974	RbCl–UCl_4	42	370.0	4446			
1975	$BeCl_2$–TlCl	18	370.0	512			
1976	$CdCl_2$–$PbCl_2$	37	370.0	870			
1977	$PbCl_2$–TlCl	40	370.0	711	763		
1978	$CdCl_2$–CsCl–TlBr	62.6–13–24.4	370.0	2562			
1979	NaCl–NaI–PbI_2	8.5–21.3–70.2	370.0	323			
1980	$CdCl_2$–$CdSO_4$–KCl	25.0–15.8–59.1	370.0	348			
1981	$CdBr_2$–NaBr	47	370.0	2071			
1982	LiBr–Li_2CrO_4	55	370.0	1938			
1983	$K_2Cr_2O_7$–$Rb_2Cr_2O_7$	50	370.0	2262			
1984	KBr–K_2CrO_4–Li_2CrO_4	15.7–18.9–65.4	370.0	2989			
1985	$CdCl_2$–CsBr–TlBr	66.7–8.3–25.0	370.0	3004			
1986	$BeCl_2$–TlCl	18	370.0	3128			
1987	KOH–K_2SO_4	94 APP	370.0	3198			
1988	$CaCl_2$–KCl–NaCl–$PbCl_2$	18.1–5.8–23.3–52.8	371.0 ±1	932			
1989	$FeCl_2$–KCl–$NdCl_3$	42–38–20	371.0	2497			
1990	$FeCl_2$–KCl–$NdCl_3$	38–42–20	372.0	2497			
1991	AgCl–$InCl_3$	40 APP	372.0	992			
1992	$BaCl_2$–$BeCl_2$	13	372.0	512			
1993	CsCl–CsI–$PbCl_2$	11.4–61.6–27	372.0	2198			
1994	K_2CO_3–Li_2CO_3–LiOH	16.4–26.4–57.1	372.0	2526			
1995	$BaCl_2$–$BeCl_2$	13	372.0	3128			
1996	$BaCl_2$–$CeCl_3$–NaCl	22–36–42	373.0	743	2447		
1997	NaCl–$PbCl_2$–TlCl	2.4–17.9–79.8	373.0	2472			
1998	$PbCl_2$–TlCl	14.9	373.0	512			
1999	$CdBr_2$–CsBr	61.9	373.0	1008	1010		
2000	K_2CrO_4–Li_2CrO_4–LiOH	19–63.5–17.5	373.0	942			
2001	$FeCl_2$–NaCl	44	374.0	1896			
2002	$CdCl_2$–CsCl–TlCl	65.3–6.6–28.1	374.0	2562			
2003	$InCl_3$–$PbCl_2$	28	374.0	397			
2004	CsCl–UCl_4	41	374.5	3246			
2005	CsCl–$NbOCl_3$	28 APP	375.0	963			
2006	$CsNO_2$–$CsNO_3$	45	375.0	1192			
2007	$TeCl_4$–TlCl	12	375.0	2707			

Locator number	System	Mol %	T, °C	References
2008	$AgVO_3$–K_2SO_4–KVO_3	11–4–85	375.0	2878
2009	KF–SnF_4	52.8	375.0	2896
2010	KCl–$ThCl_4$	55.5	376.0	1922
2011	CsI–$PbCl_2$	71.8	376.0	2198
2012	$CdBr_2$–TlBr	55	376.0	2071
2013	Na_2SO_4–Tl_2SO_4–$TlVO_3$	0.5–0.6–98.9	376.0	2983
2014	$NaHSO_4$–NH_4HSO_4	29	376.2	2455
2015	AgCl–Ag_2S	64.2	377.0	989
2016	AgCl–Ag_2Te	80	377.0	2829
2017	NaCl–UCl_4	52	377.9	3246
2018	$DyCl_3$–NaCl	55	378.0	1046
2019	CsCl–$GaCl_3$	58	378.0	1016
2020	CuCl–$CuCl_2$	87	378.0	38
2021	$BaCl_2$–$CaCl_2$–LiCl–NaCl	15.9–34.5–29.1–20.5	378.0	2963
2022	Na_2SO_4–$NaVO_3$–$TlVO_3$	0.3–7.4–92.3	378.0	2983
2023	KCl–$ThCl_4$–UCl_3	30.5–54–15.5	378.0 ±2	2805
2024	K_2CrO_4–KF–Li_2CrO_4	4–31–65	378.0	2855
2025	AgF–ZnF_2	86	380.0	536
2026	$CdCl_2$–CsCl–NaCl	58.7–2.4–38.9	380.0	320
2027	KCl–NaCl–$SmCl_3$	27.1–25.6–47.3	380.0	1186
2028	CsCl–KCl–TlCl	25–7–68	380.0	2257
2029	$FeCl_2$–KCl	52.2	380.0	2497
2030	KCl–$MgCl_2$–$TiCl_3$	45.2–47.4–7.4	380.0	434
2031	KCl–$MgCl_2$–$TiCl_3$	55.6–33.2–11.2	380.0	434
2032	KCl–$NbOCl_3$	33.2	380.0	1800
2033	$PbBr_2$–$PbCl_2$	36	380.0	141 167 433 511 738
2034	NaCl–PbI_2	15.7	380.0	323
2035	AgCl–Ag_2S	65	380.0	717
2036	KBr–$PbBr_2$–TlBr	12–18–70	380.0	782
2037	$CsNO_3$–TlBr	85.2	380.0	1170
2038	$Cs_2O(Cs_2CO_3)$–V_2O_5	39	380.0	854
2039	Rb_2O–V_2O_5	39.5	380.0	2069
2040	Rb_2O–V_2O_5	41	380.0	1134
2041	Na_2SO_4–$TlVO_3$	1.0	380.0	2983
2042	LiCl–$ThCl_4$–UCl_4	50.0–8.0–42.0	380.0 ±2	3231
2043	AgBr–AgI	79 APP	380.0	3167
2044	CsCl–LiCl	58	381.0	2759
2045	NaCl–$TbCl_3$	55	382.0	1482
2046	CsCl–PbI_2	54.5	382.0	2198
2047	CsCl–KCl–$PbCl_2$	7–21–72	382.0	3234
2048	$CaCl_2$–NaCl–$PbCl_2$	19–32–49 APP	382.0	2660
2049	$Li_2Cr_2O_7$–$Na_2Cr_2O_7$	70 APP	382.0	2924
2050	$PbCl_2$–RbCl–TlCl	10–10–80	382.0	2833
2051	PbF_2–PbI_2	10	383.0	802
2052	NaCl–$NbOCl_3$	33.4	383.0	1800
2053	$CdCl_2$–KCl	66.7	383.0	13 104 259 322 348 394 409 498
2054	CsBr–TlCl	28	383.0	2469
2055	Li_2CrO_4–Na_2CrO_4	68.5	383.0	3032
2056	Li_2CrO_4–Na_2CrO_4	68.5	383.0	3156
2057	$CdCl_2$–NaCl–TlCl	39–23–38	384.0	480
2058	NaCl–$PbCl_2$–TlCl	14.9–66–19.1	384.0	2472
2059	KCl–$ThCl_4$	57.2	384.0	54
2060	$CdBr_2$–CsBr	64	384.0	2071
2061	K_2SO_4–Na_2SO_4–$ZnSO_4$	30.2–29.1–40.7	384.0	212
2062	$NaBF_4$–NaF	92±1	384.0 ±2	2703
2063	AgCl–Tl_2SO_4	86	384.0	3124
2064	$NaBF_4$–NaF	92	384.0	1039

TABLE 1. Eutectic data—Continued

Locator number	System	Mol %	T, °C	References
2065	$CdCl_2$–$PbCl_2$	36.5	385.0	41 61 62 322 394 786 1676
2066	CsI–$PbCl_2$	15.7	385.0	2198
2067	$CdBr_2$–CdI_2	8 APP	385.0	1676
2068	$CdBr_2$–CsI	26.2	385.0	1010
2069	CuBr–CuI	82.5	385.0	1918
2070	PbI_2–PbTe	99	385.0	1971
2071	K_2SO_4–Na_2SO_4–$ZnSO_4$	19.0–25.8–55.2	385.0	212
2072	CsCl–KCl–$PbCl_2$	9–4–87	385.0	3234
2073	$CsNO_2$–$CsNO_3$	50 APP	385.0 MIN	2923
2074	$BaCl_2$–$CaCl_2$–LiCl–NaCl	16.6–37.6–33.3–12.5	385.0	2963
2075	$CdCl_2$–$PbCl_2$	30	385.0	3138
2076	CsCl–PbI_2	77.7	386.0	2198
2077	TlBr–Tl_2SO_4	73 APP	386.0	3216
2078	$CdCl_2$–KCl–LiCl	13.6–18.2–68.2	387.0	880
2079	$CdCl_2$–NaCl	55	387.0	304 321 480
2080	KCl–NaCl–$TaCl_5$	49.1–1.8–49.1	388.0	797
2081	$CdCl_2$–KCl	38	388.0	13 104 259 322 348 394 409 498
2082	$CdBr_2$–TlCl	59.4	388.0	2297
2083	LiCl–$PbCrO_4$	72.6	388.0	1054
2084	$CsNO_3$–TlI	90.9	388.0	1170
2085	K_2SO_4–Na_2SO_4–$ZnSO_4$	24.5–33.4–42.1	388.0	212
2086	$CdCl_2$–$PbCl_2$	35	389.0	41 61 62 322 394 786
2087	KCl–$MnCl_2$–NaCl	45.5–34.5–20	390.0	2524
2088	KCl–NaCl–$NbCl_5$	49–2–49	390.0	1118
2089	NaCl–$SmCl_3$	51	390.0	1011
2090	$FeCl_2$–KCl	53	390.0 ±2	1896
2091	CsCl–TlCl	25	390.0	512
2092	$InCl_2$–TlCl	13 APP	390.0 APP	1462
2093	$InCl_3$–TlCl	6.4	390.0	873
2094	$CdCl_2$–$CdSO_4$–TlCl	63.3–3.3–33.3	390.0	392
2095	AgCl–Ag_2WO_4	82.3	390.0	765
2096	LiCl–$ThCl_4$–UCl_4	59.0–19.0–22.0	390.0 ±2	3231
2097	LiCl–$ThCl_4$–UCl_3	51–34–15	390.0	2827
2098	AgI–NaI	75	390.0	3115
2099	LiH–LiI	23.5	390.8	1321
2100	$CaCl_2$–NaCl–$PbCl_2$	17.8–25.2–57	391.0	512
2101	CuCN–NaCN	NA	391.0	3111
2102	$CdCl_2$–NaCl	55	392.0	786
2103	LiCl–TlBr	23 APP	392.0	887
2104	CuCl–Cu_2S	89.2	392.0	805
2105	LiCl–Li_3VO_4–$PbCl_2$	23.5–4.8–71.6	392.0	523
2106	NaCl–$PbCl_2$–PbS	24.6–68–7.4	392.0	733
2107	$CdBr_2$–TlBr	59	392.0	788
2108	K_2SO_4–Na_2SO_4–$ZnSO_4$	34.0–22.5–43.5	392.0	212
2109	BeF_2–CsF	58.4	393.0	1986
2110	$FeCl_2$–KCl	54.2	393.0 ±1	574
2111	$CaCl_2$–CuCl	10.8	393.0	156
2112	CsI–TlCl	61	393.0	2240
2113	$Ba(ClO_4)_2$–$KClO_4$	71	393.0	1116
2114	$Ca(NO_2)_2$–$Ca(NO_3)_2$	94	393.0	915
2115	K_2CO_3–Li_2CO_3–Na_2CO_3	26.8–42.5–30.6	393.0	1137
2116	K_2CrO_4–$K_2Cr_2O_7$	1 APP	393.0	3189
2117	NaCl–YCl_3	55	394.0	2236
2118	$CdBr_2$–CsBr–KBr	25–56.2–18.7	394.0	2527
2119	$CaCl_2$–NaCl–YCl_3	5–51–44	395.0	1154
2120	KBr–$ZnSO_4$	67.5	395.0	3249